Dr. Nervana Mostafa

MB BS, MD, PhD (UK)

Professor of Physiology
College of Medicine, KKUH, KSU

nbayoumy@ksu.edu.sa

Cell membrane structure & transport across cell membrane

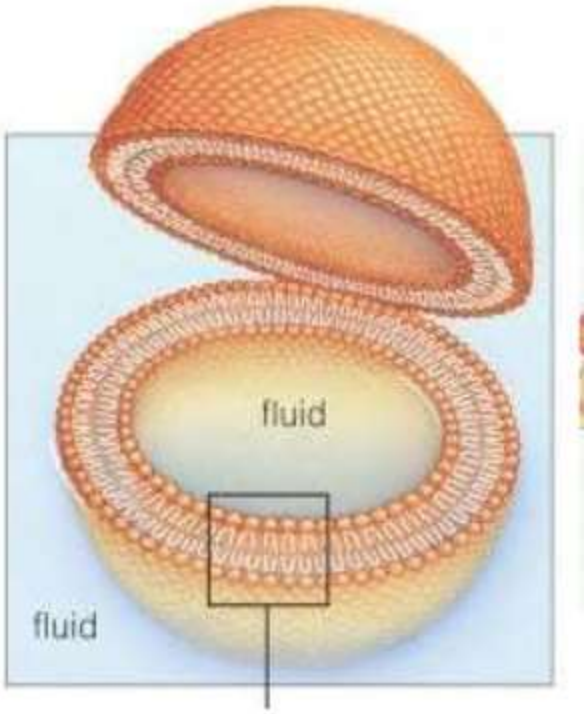
objectives

- **Describe the fluid mosaic model of membrane structure and function.**
- **Define permeability and list factors influencing permeability.**
- **Identify and describe carried-mediated transport processes: Primary active transport, secondary active transport, facilitates diffusion.**
- **Differentiate between passive and active transport mechanisms and give examples on each.**

#Study source for this lecture: (Guyton & Hall Textbook of Medical Physiology, 13th edition) #

Cell Membrane

- Envelops the cell.
- Thin, pliable and elastic.
- 7 - 10 nanometer thick.



- Also, referred to as the plasma membrane .

Composition

Lipoprotein

protein 55%

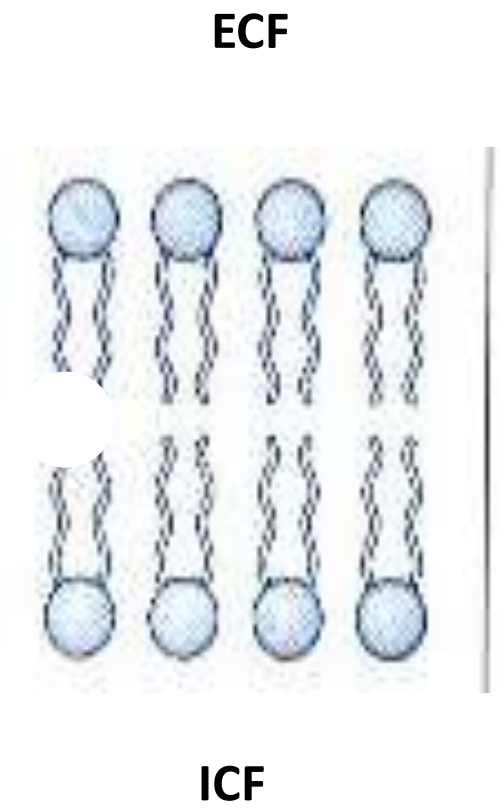
lipid 42%
- phospholipids 25%
- cholesterol 13%
- glycolipid 4%

carbohydrates 3%

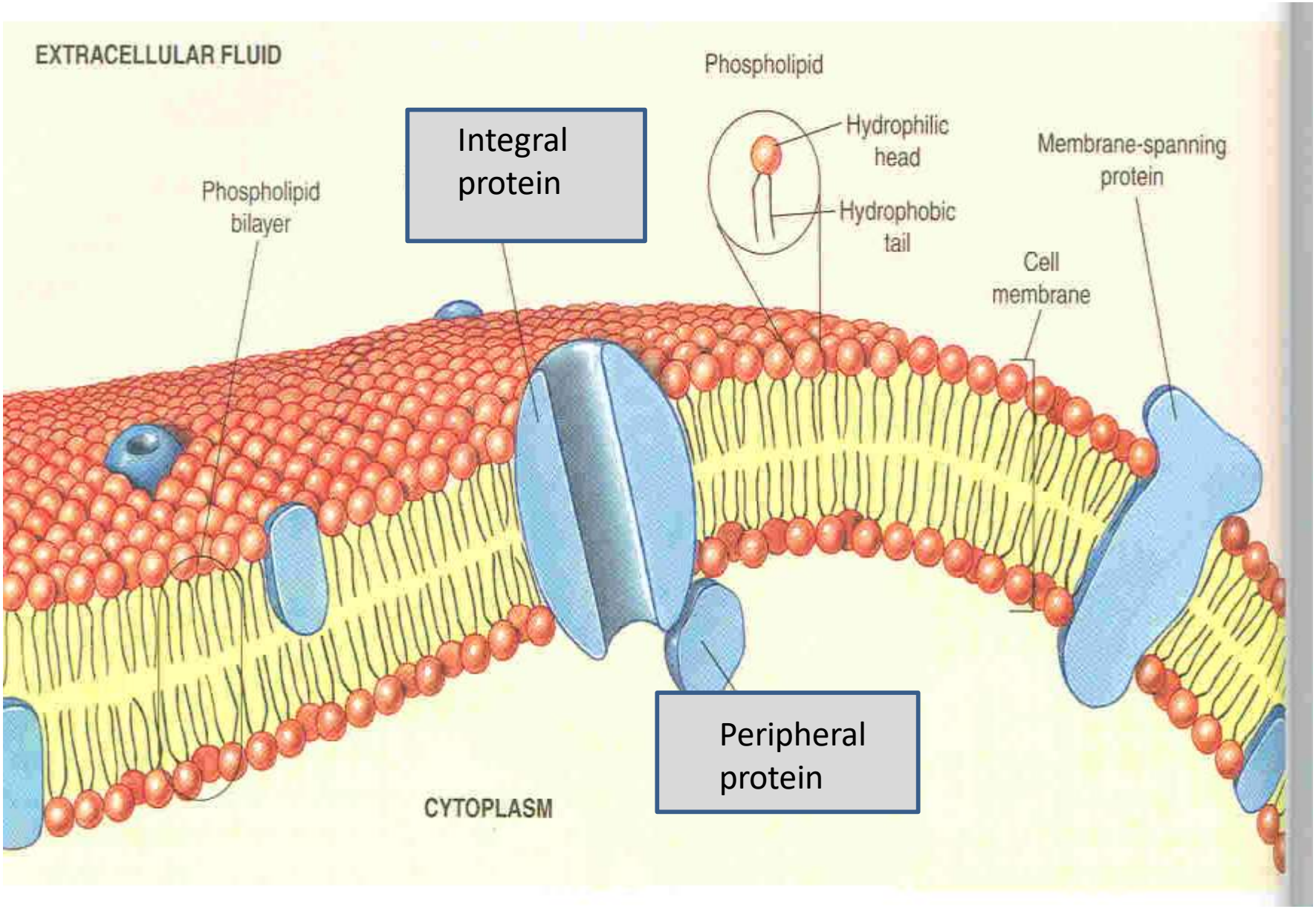
The Cell Membrane Phospholipids Consist Of :

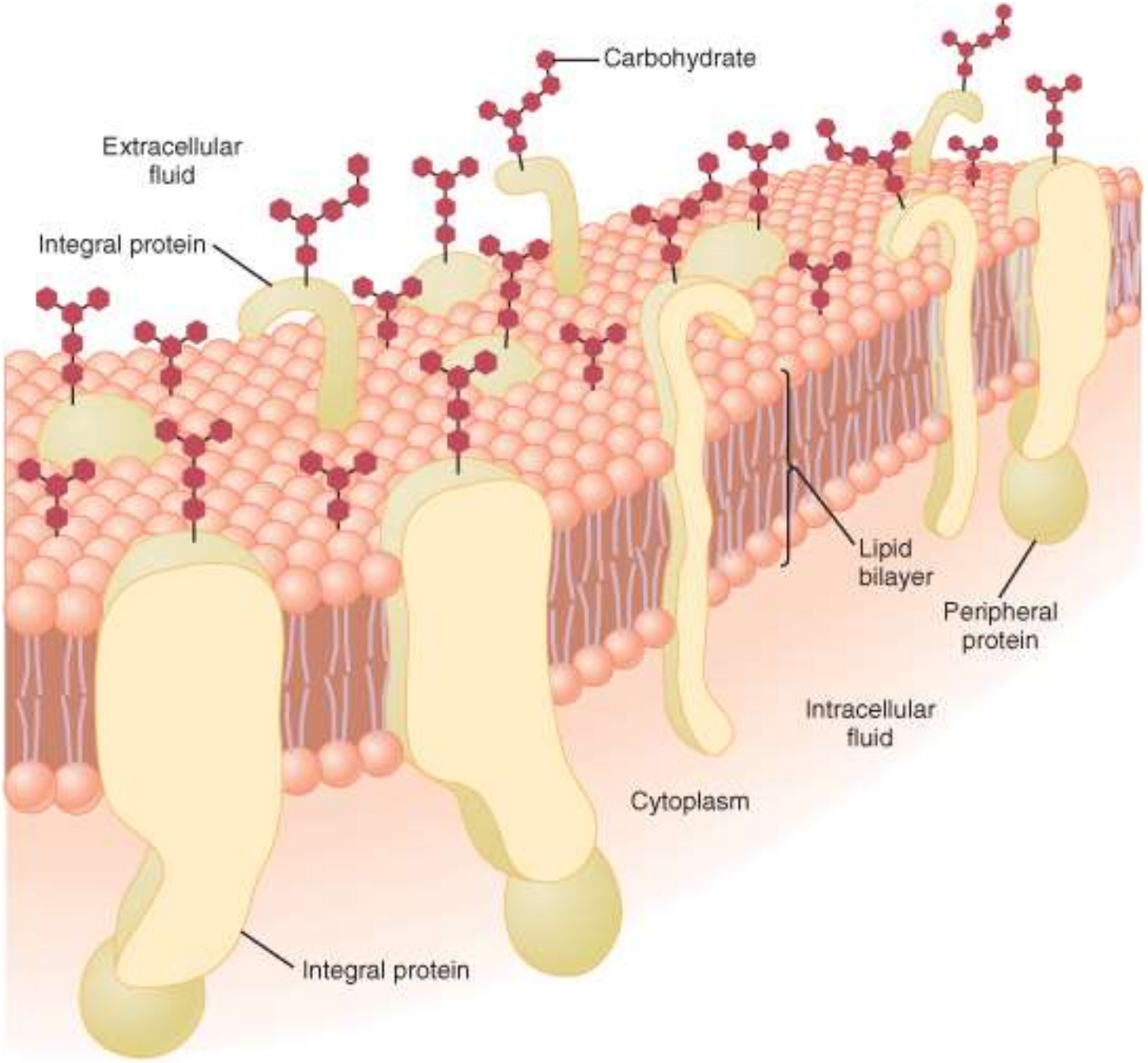
1. **Glycerol head** (hydrophilic).

2. **Two fatty acid ‘’tails’’** (hydrophobic).



The Cell Membrane Proteins





The Cell Membrane Proteins:

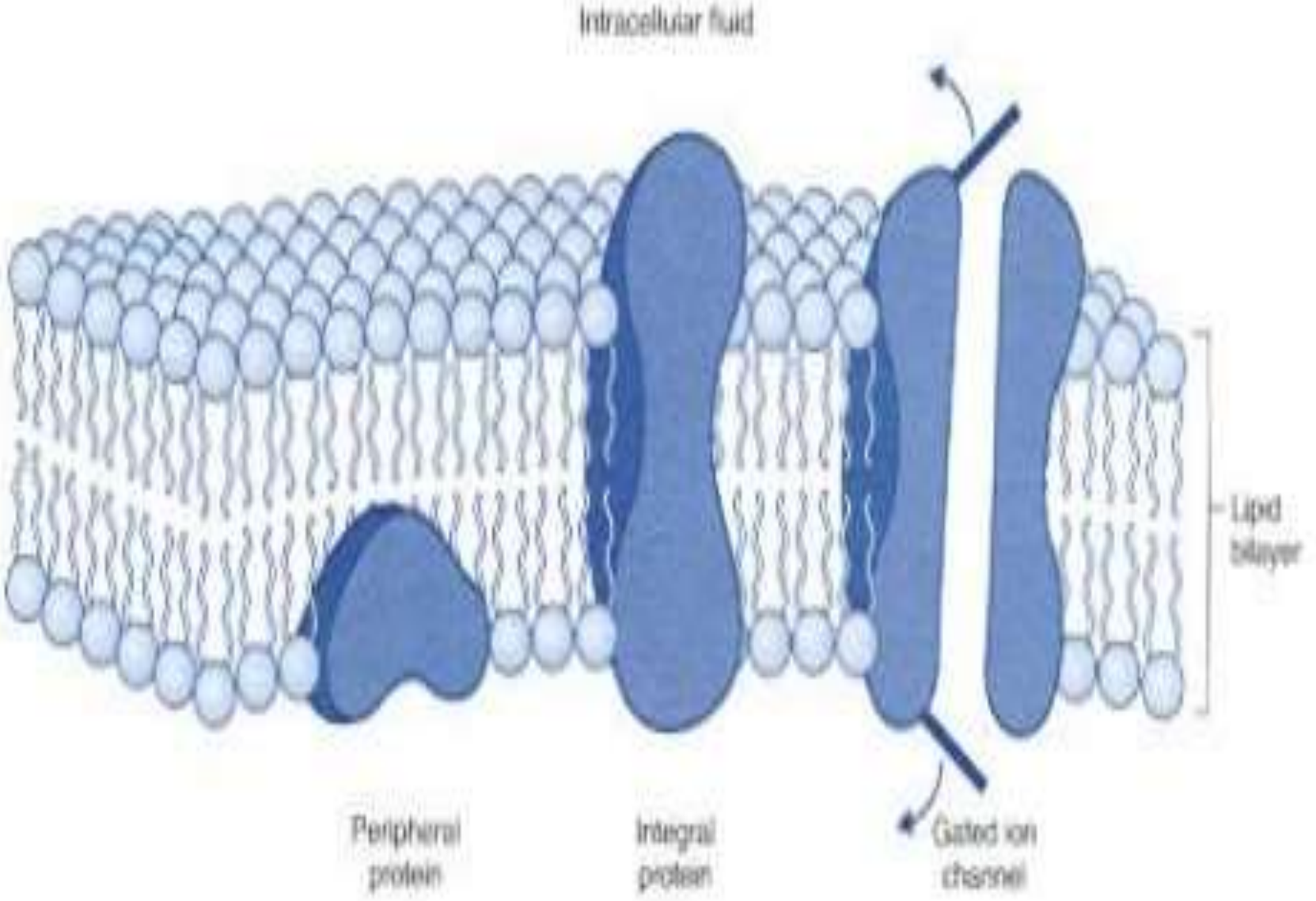
1. Integral proteins

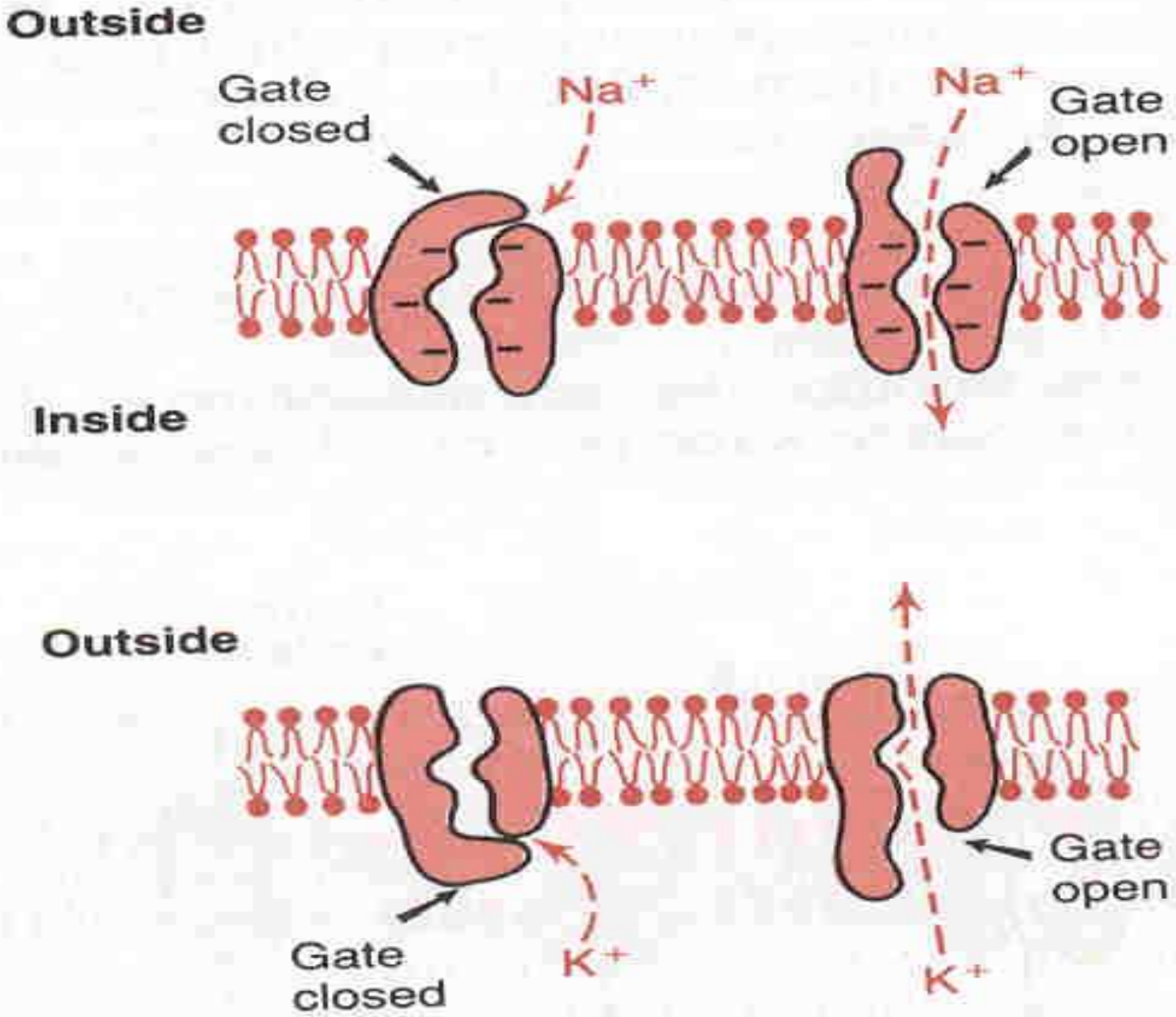
- Span the whole thickness of the membrane .
- Proteins provide structural **channels** or **pores**.
- **Carrier** proteins.

2. Peripheral proteins

-Present in one side.

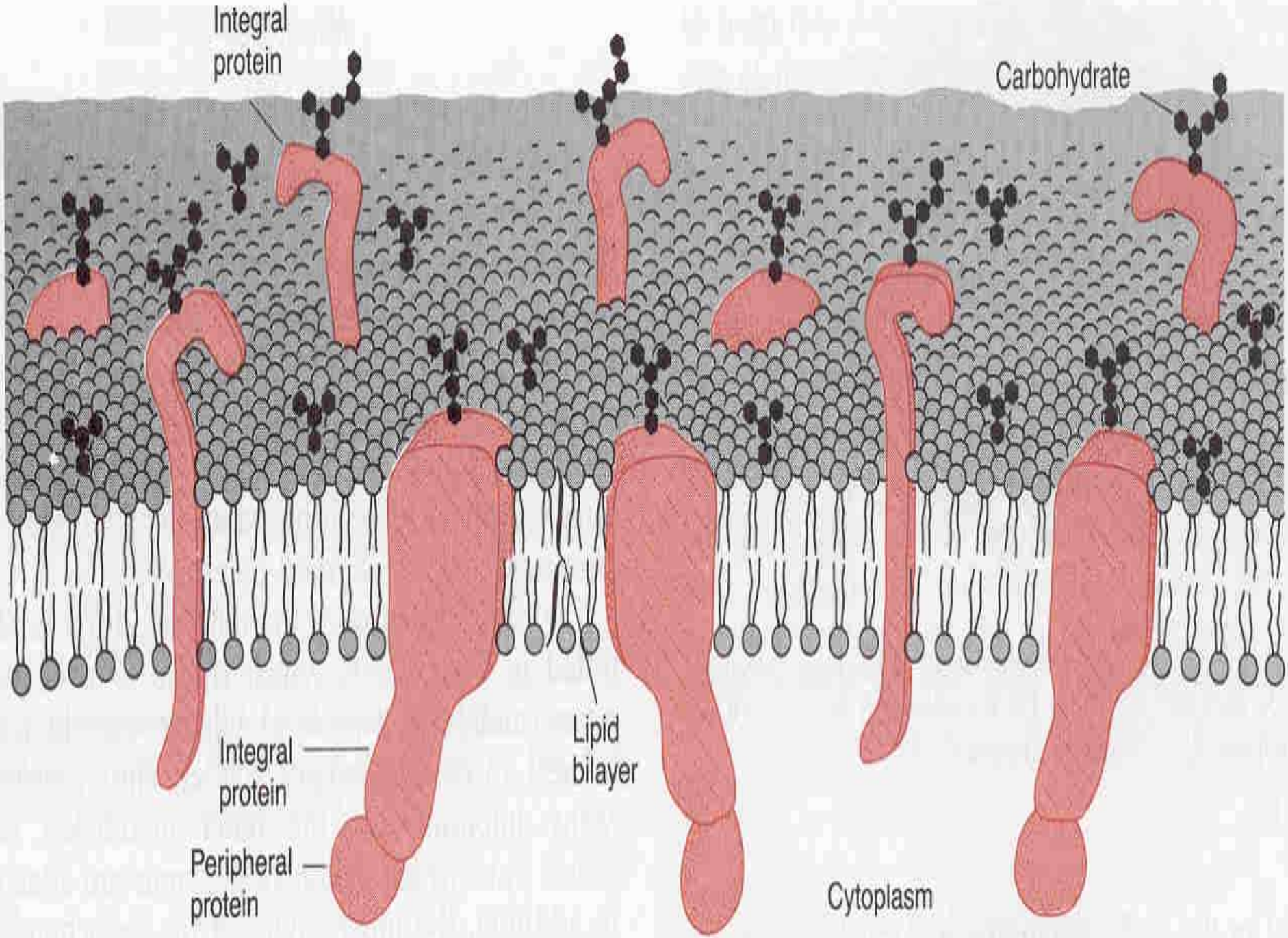
- **Hormone receptors .**
- **Cell surface antigens .**


Intracellular fluid
Lipid bilayer
Peripheral protein
Integral protein
Gated ion channel
Extracellular fluid


Outside
Gate closed
Na+
Na+
Gate open
Inside
Outside
Gate closed
K+
K+
Gate open
Inside

The Cell Membrane Carbohydrates:

- **Glycoproteins** (most of it) and **Glycolipids**.
- **Proteoglycans** (mainly carbohydrate substance bound together by protein).
- ‘’glyco’’ part is in the surface (hydrophilic) .
- Glycocalyx (Carbohydrate molecules protrude to the outside of the cell forming a loose carbohydrate coat “glycocalyx”.


Integral protein
Carbohydrate
Integral protein
Lipid bilayer
Peripheral protein
Cytoplasm

Function Of Carbohydrates:

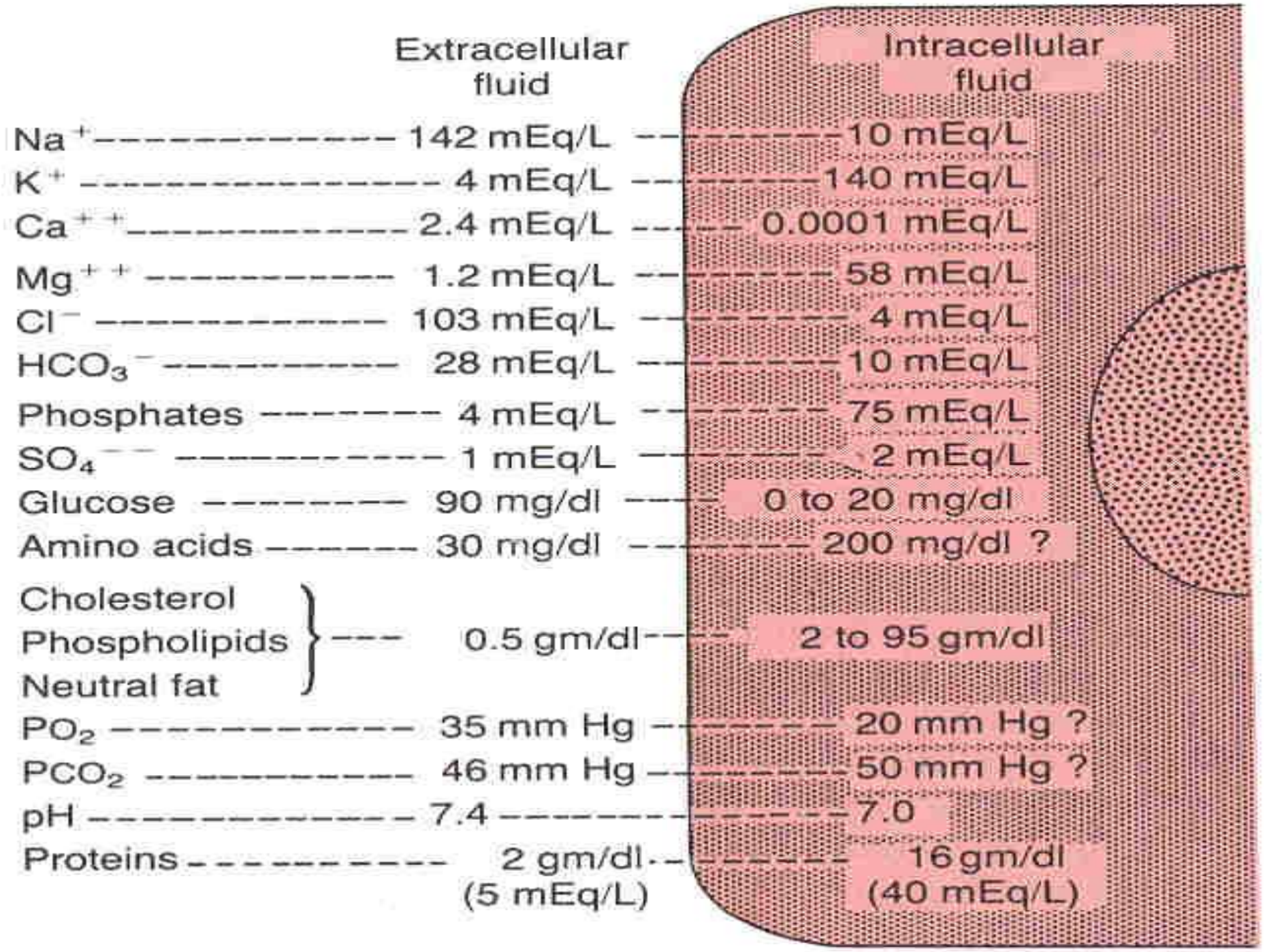
- **Attaches** cell to each others.
- Act as **receptors** substances (help ligend to recognize its receptor).
- Some enter in to **immune reactions**.
- Give most of cells overall **–ve surface**.

Transport across the Cell Membrane

- Cell membrane is ***selectively permeable***.

- Through the **proteins**.
  - Water-soluble substances e.g. ions, glucose ..

- Directly through the lipid **bilayer**.
  - Fat-soluble substance (O2, CO2, N2, alcohol..


Extracellular fluid
Intracellular fluid
Na^+ 142 mEq/L 10 mEq/L
K^+ 4 mEq/L 140 mEq/L
Ca^{++} 2.4 mEq/L 0.0001 mEq/L
Mg^{++} 1.2 mEq/L 58 mEq/L
Cl^- 103 mEq/L 4 mEq/L
HCO_3^- 28 mEq/L 10 mEq/L
Phosphates 4 mEq/L 75 mEq/L
SO_4^{--} 1 mEq/L 2 mEq/L
Glucose 90 mg/dl 0 to 20 mg/dl
Amino acids 30 mg/dl 200 mg/dl ?
Cholesterol
Phospholipids
Neutral fat
0.5 gm/dl
2 to 95 gm/dl
PO_2 35 mm Hg 20 mm Hg ?
PCO_2 46 mm Hg 50 mm Hg ?
pH 7.4 7.0
Proteins 2 gm/dl (5 mEq/L) 16 gm/dl (40 mEq/L)

Types Of Membrane Transport

1- Diffusion

a) simple diffusion.

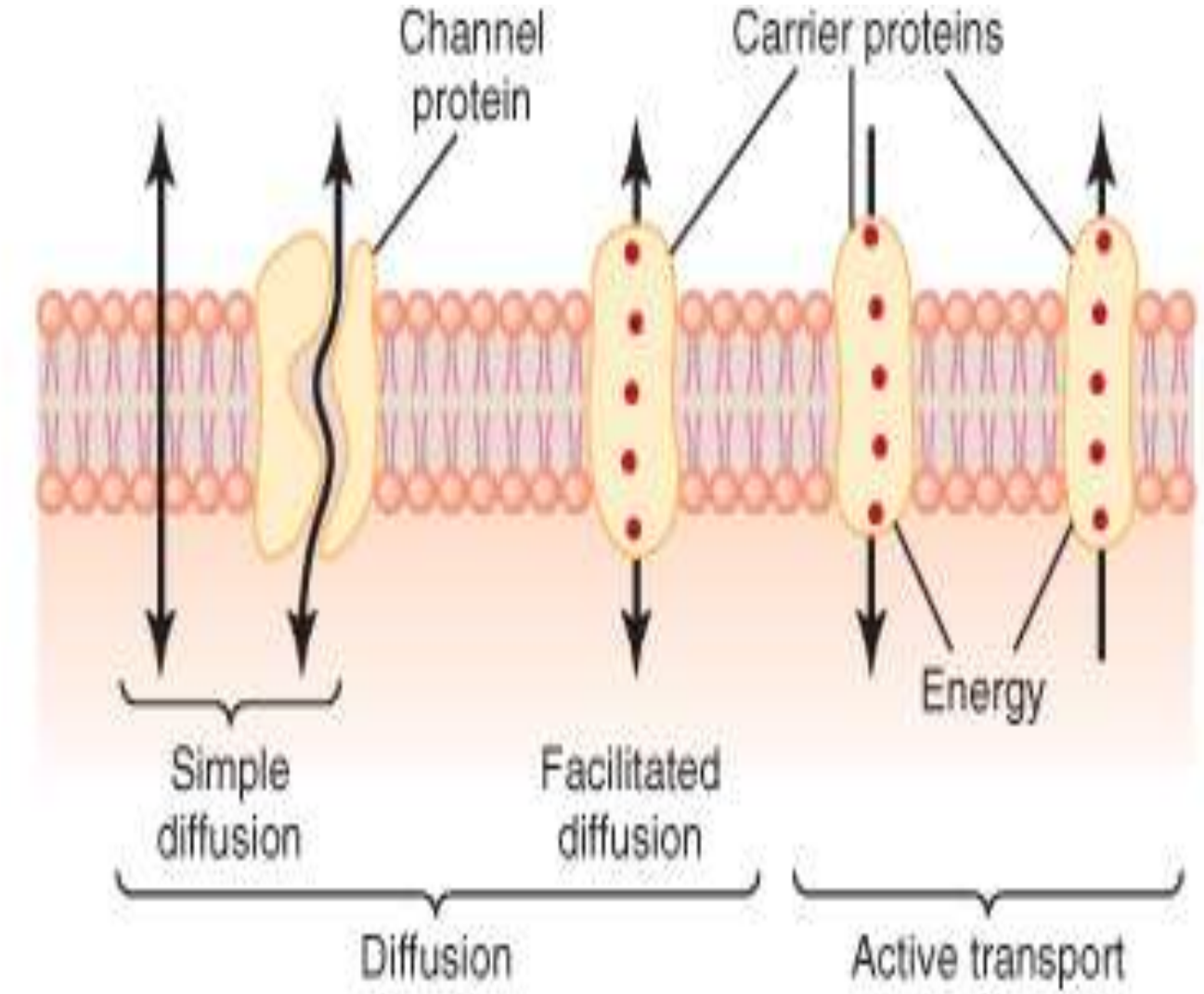
b) facilitated diffusion.

2- Active transport.

a) primary active transport.

b) secondary active transport.

3- Osmosis.



Diffusion

Random movement of substance either through the membrane directly or in combination with carrier protein down an electrochemical gradient.

1- Simple diffusion.

2- facilitated diffusion.

- Simple and facilitated diffusion Do NOT require input of energy = powered by concentration gradient or electrical gradient.
- Active transport = uses energy = utilizes ATP.

Simple Diffusion

- Non-carrier mediated transport down an electrochemical gradient.
- Diffusion of non-electrolytes (uncharged) from high concentration to low concentration.
- Diffusion of electrolytes (charged) depend on both chemical, as will as, electrical potential difference.

Rate Of Simple Diffusion Depend On:

1- Amount of substance available.

2- The number of opening in the cell membrane for the substance (pores).

3- Chemical concentration difference.

4- Electrical potential difference.

5- Molecular size of the substance.

6- Lipid solubility.

7- Temperature.

TABLE 3-1 Factors Influencing the Rate of Net Diffusion of a Substance across a Membrane (Fick's Law of Diffusion)

Factor	Effect on Rate of Net Diffusion
↑ Concentration gradient of substance (ΔC)	↑
↑ Surface area of membrane (A)	↑
↑ Membrane permeability	↑
↑ Molecular weight of substance (MW)	↓
↑ Distance (thickness) (ΔX)	↓
↑ Temperature	↑

Factors Affecting Rate of Diffusion *cont…*

$$Rate\ of\ diffusion = P\ X\ A\ (C1–C2)$$

1. **P = Permeability coefficient.**

a. Temperature **b. Size of molecule**

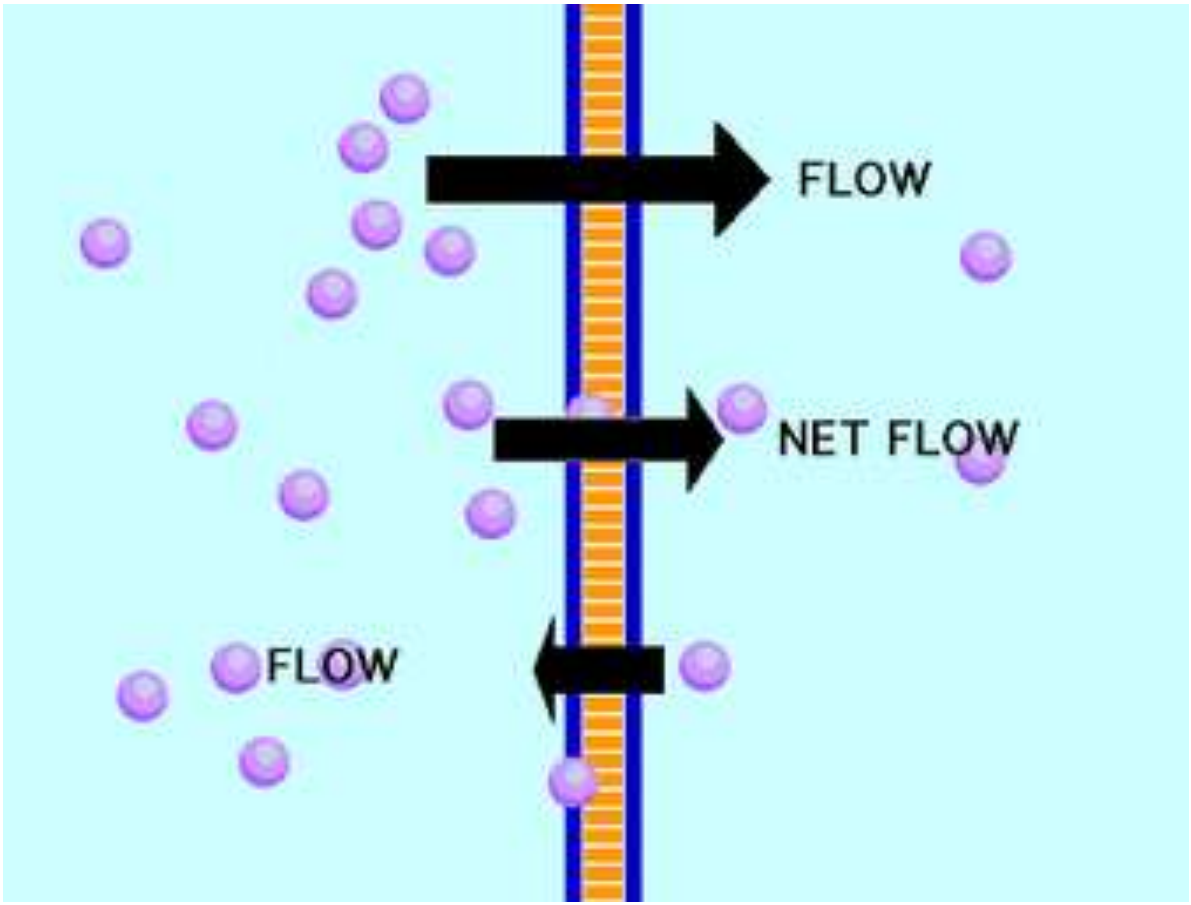
c. Solubility in lipids **d. Thickness of membrane**

2. A = surface area.

3. (C1-C2) = gradient difference:

a. **Concentration difference**
b. **Electrical difference.**
c. **Pressure difference.**

Net Flow

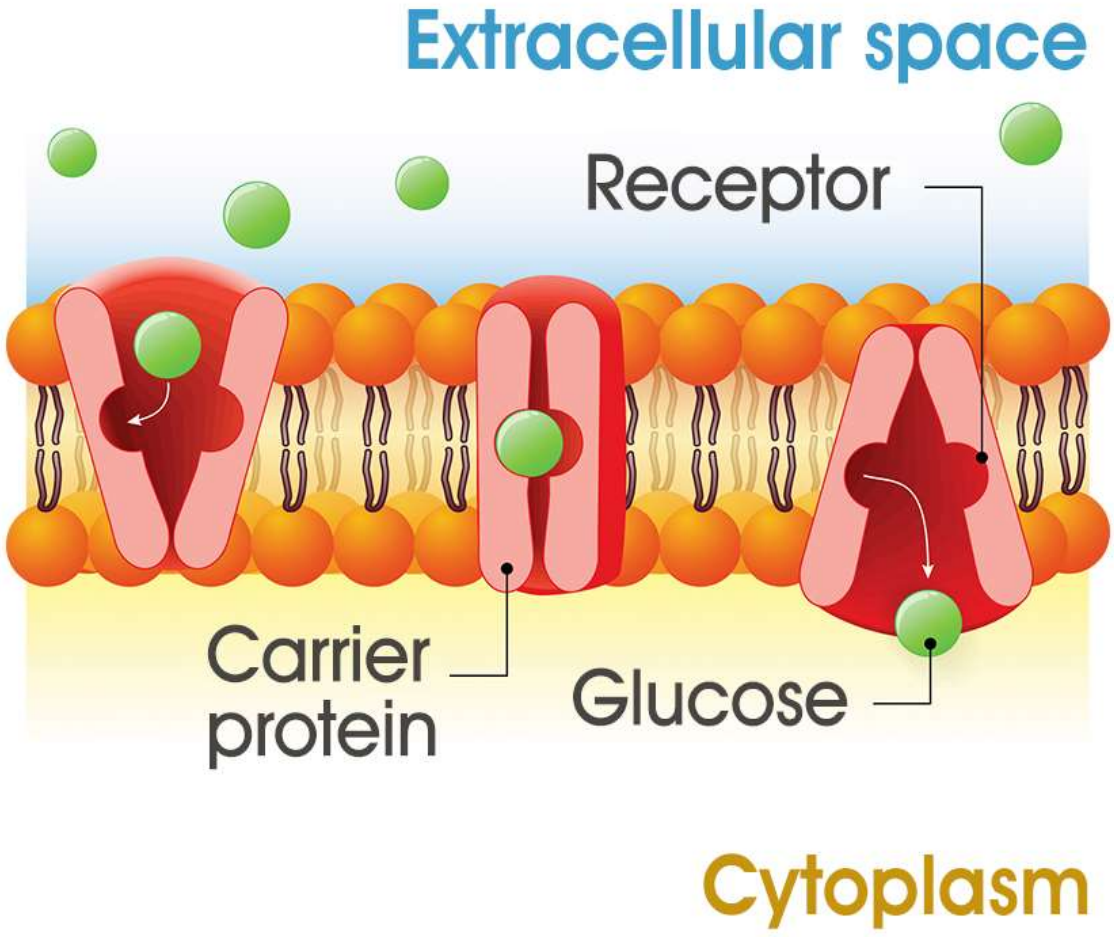


Facilitated Diffusion

- Carrier mediated transport down an electrochemical gradient.

- e.g. glucose & amino acids.

Facilitated diffusion—the process that allows selective movement in and out of the **cell membrane**.



Features Of Carrier Mediated Transport (Facilitated diffusion)

1- saturation:

↑concentration ⟶ ↑binding of protein

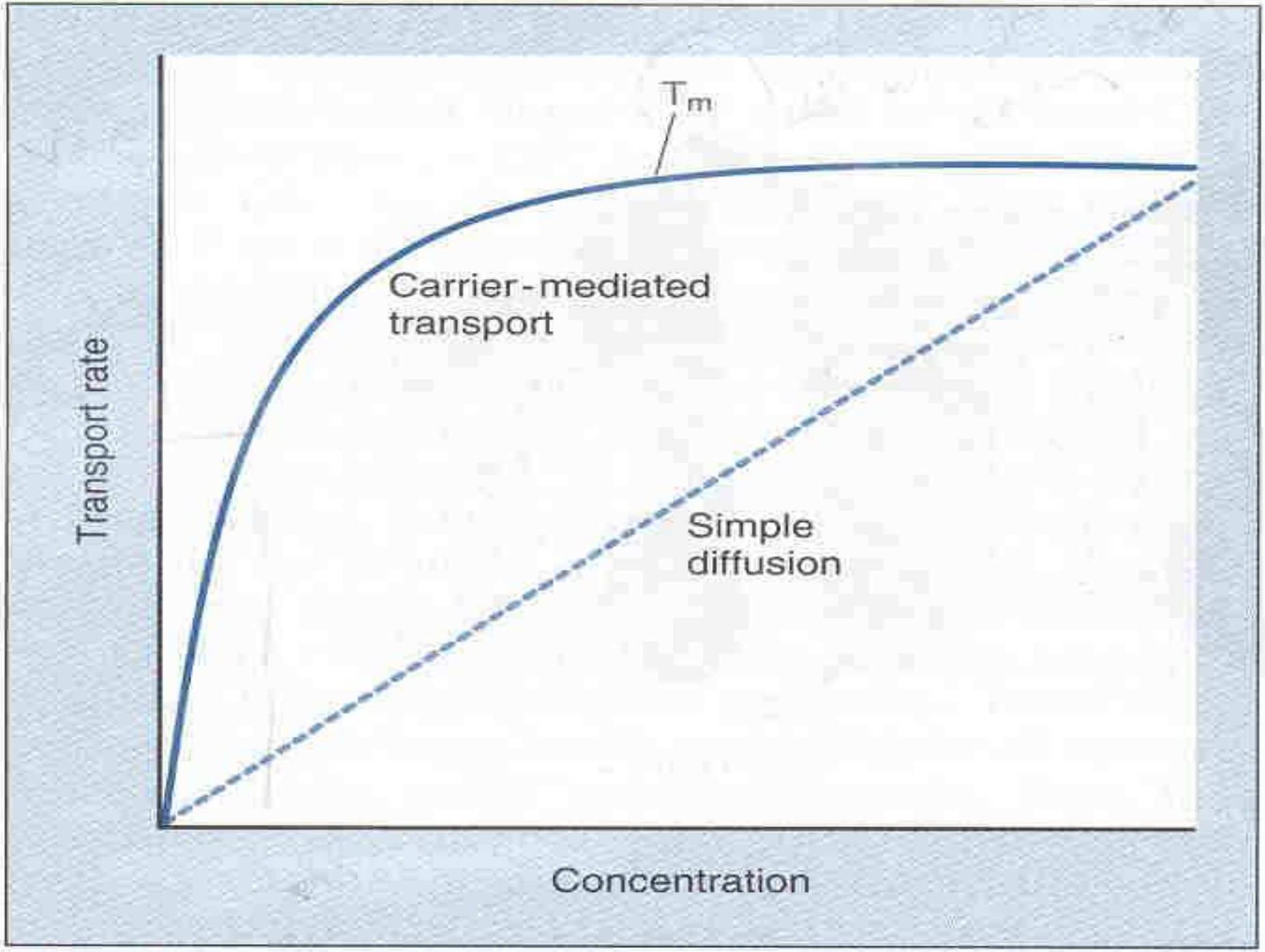
If all protein carriers are occupied, full saturation is achived.

i.e. The rate of diffusion reaches a maximum (Vmax) when all the carriers are functioning as rapidly as possible.

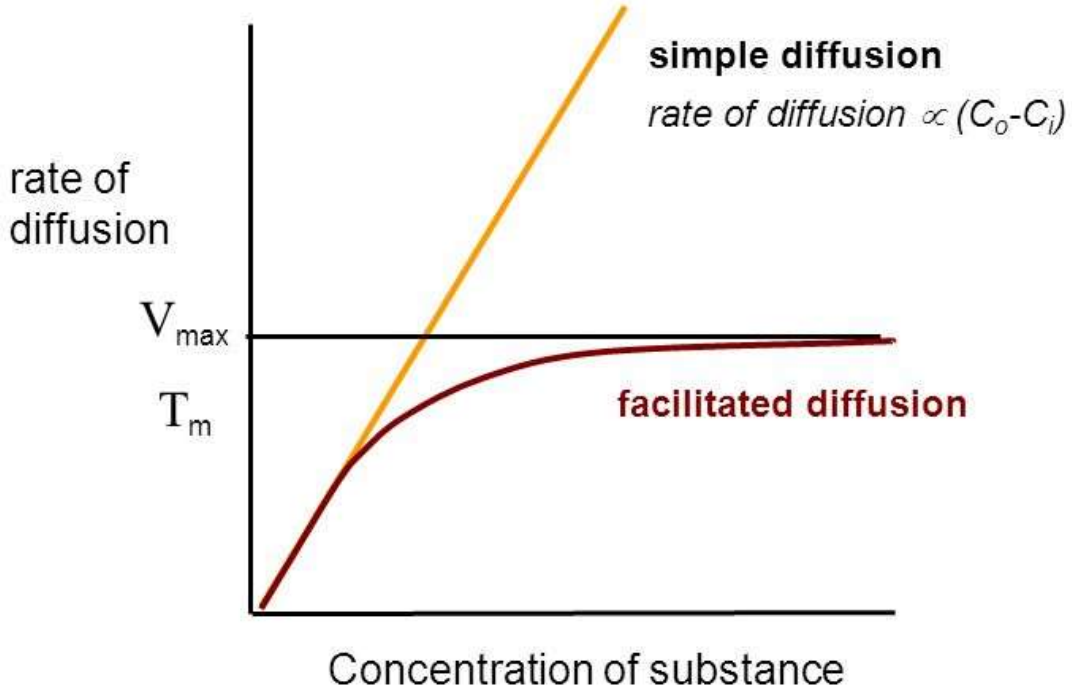
2- stereopecificity:

The binding site recognize a specific substance

D-glucose but not L-glucose.


T_m
Carrier-mediated transport
Simple diffusion
Transport rate
Concentration

Simple vs. Facilitated



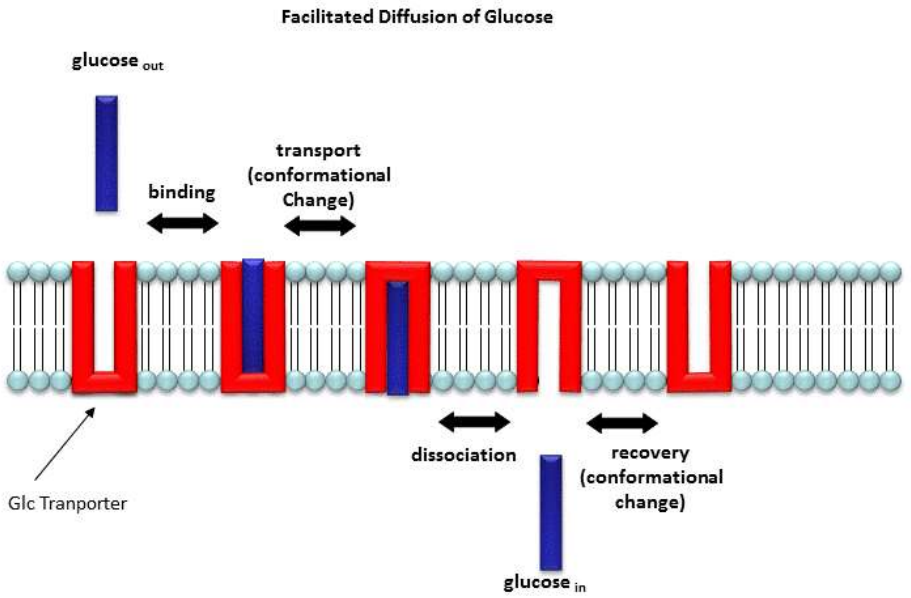
What limits maximum rate (Vmax) of facilitated diffusion? Number of carriers

3- Competition:

Chemically similar substances can compete for the same binding site.

D- galactose / D-glucose.

{Substance → binding site → substance protein complex → conformational changes → release of substance}



Active Transport:

- Transport (uphill) → against electrochemical gradient.
- Required energy → direct.
  indirect.
- Required carrier – protein.

1- Primary Active Transport:

-Energy is supplied directly from ATP.

ATP $\longrightarrow$ ADP + P + energy.

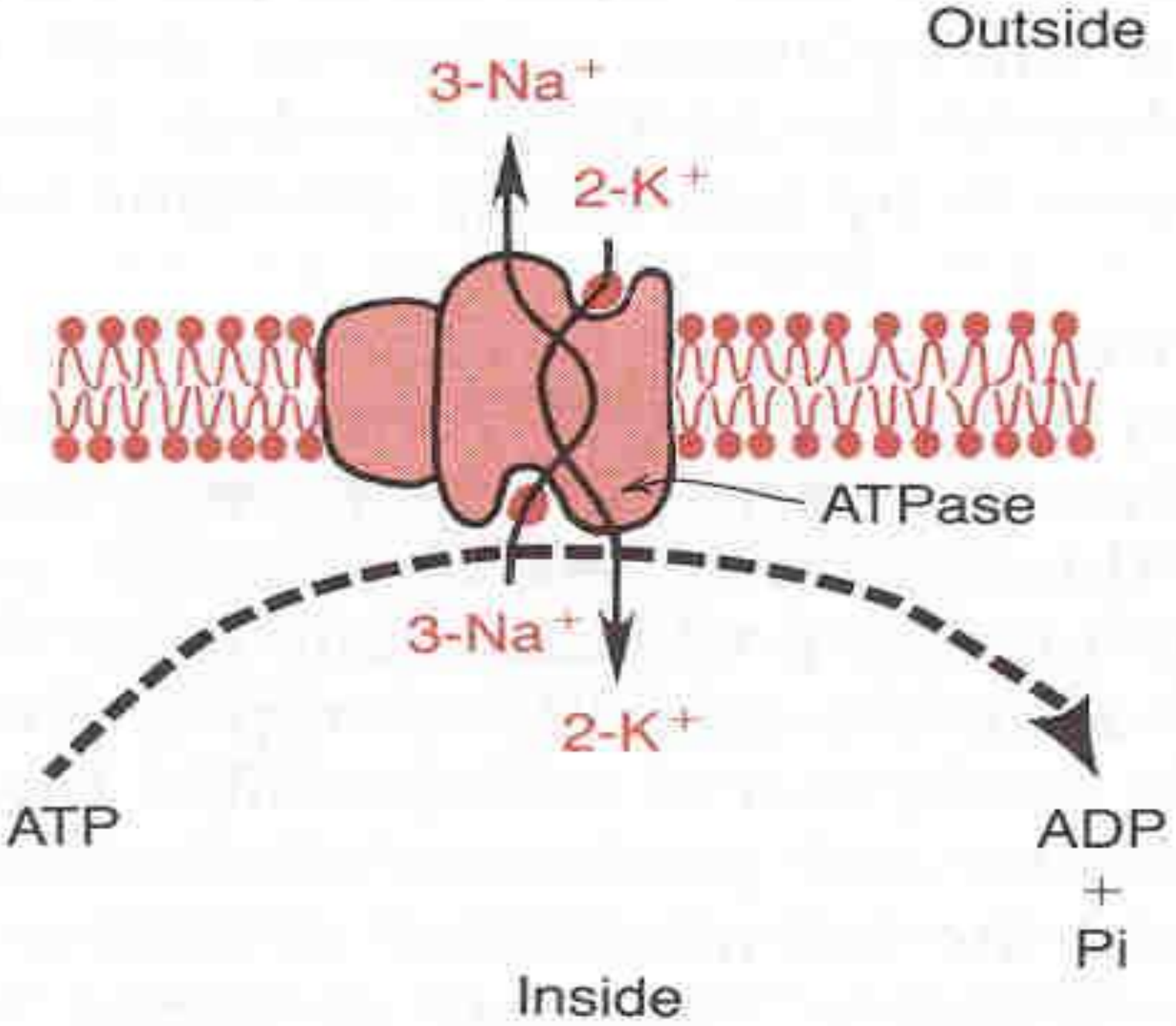
A. - **Sodium-Potassium pump (Na^+- K^+ pump).**

- it's present in all cell membranes.
- 3 Na^+ in $\longrightarrow$ out.
- 2 K^+ out $\longrightarrow$ in.

Discovery

- Na^+/K^+-ATPase pump was discovered by Jens Christian Skou in 1957.

- In 1997, he received the Nobel Prize in Chemistry .


Outside
3-Na+
2-K+
ATPase
3-Na+
2-K+
ATP
ADP
+
Pi
Inside

Characteristic of Na^+/K^+-ATPase Pump:

1. Carrier protein is formed from α and β subunits.
2. Binding site for Na inside the cell.
3. Binding site for K outside the cell.
4. It has ATPase activity.
5. 3 Na out.
6. 2 K in.

Function:

1. Maintaining Na^+ and K^+ concentration difference .

2. It’s the basis of nerve signal transmition .

3. Maintaining negative potential inside the cell.

4. Maintains a normal cell volume.

B. **primary active transport of calcium**

(Ca^{++} ATPase).

- sarcoplasmic reticulum (SR).
- mitochondria.
- in some cell membranes.

Function:

Maintaining a low Ca^{2}+ concentration inside the cell.

C. primary active transport of hydrogen ions H^+-K^+ ATPase.

- stomach.
- kidneys.
- pump to the lumen.
- H^+-K^+ ATPase inhibitors, are used to treat ulcers (e.g. omeprazole).

2) Secondary Active Transport:

- **Co- transport and counter-transport:**

- is transport of one or more solutes against an electrochemical gradient, coupled with the transport of another solute down an electrochemical gradient.

- ‘’downhill’’ solute is Na^+.

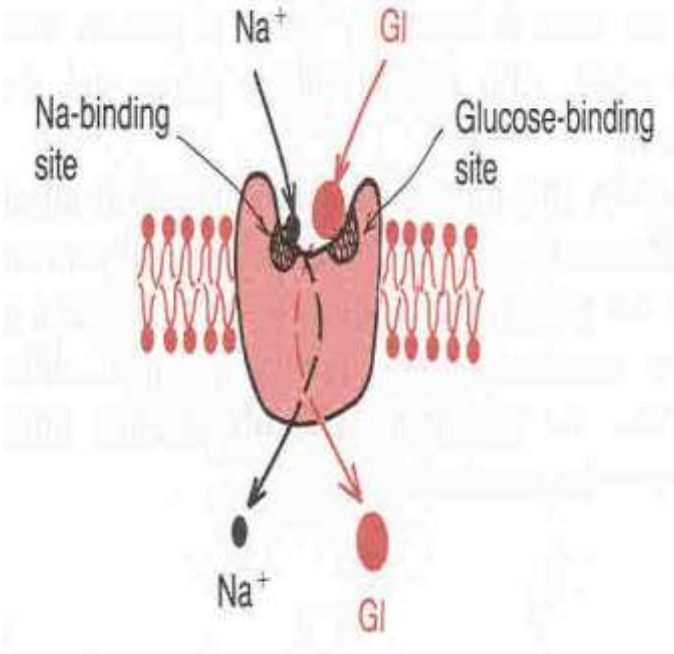
- Energy is supplied indirectly form primary transport.

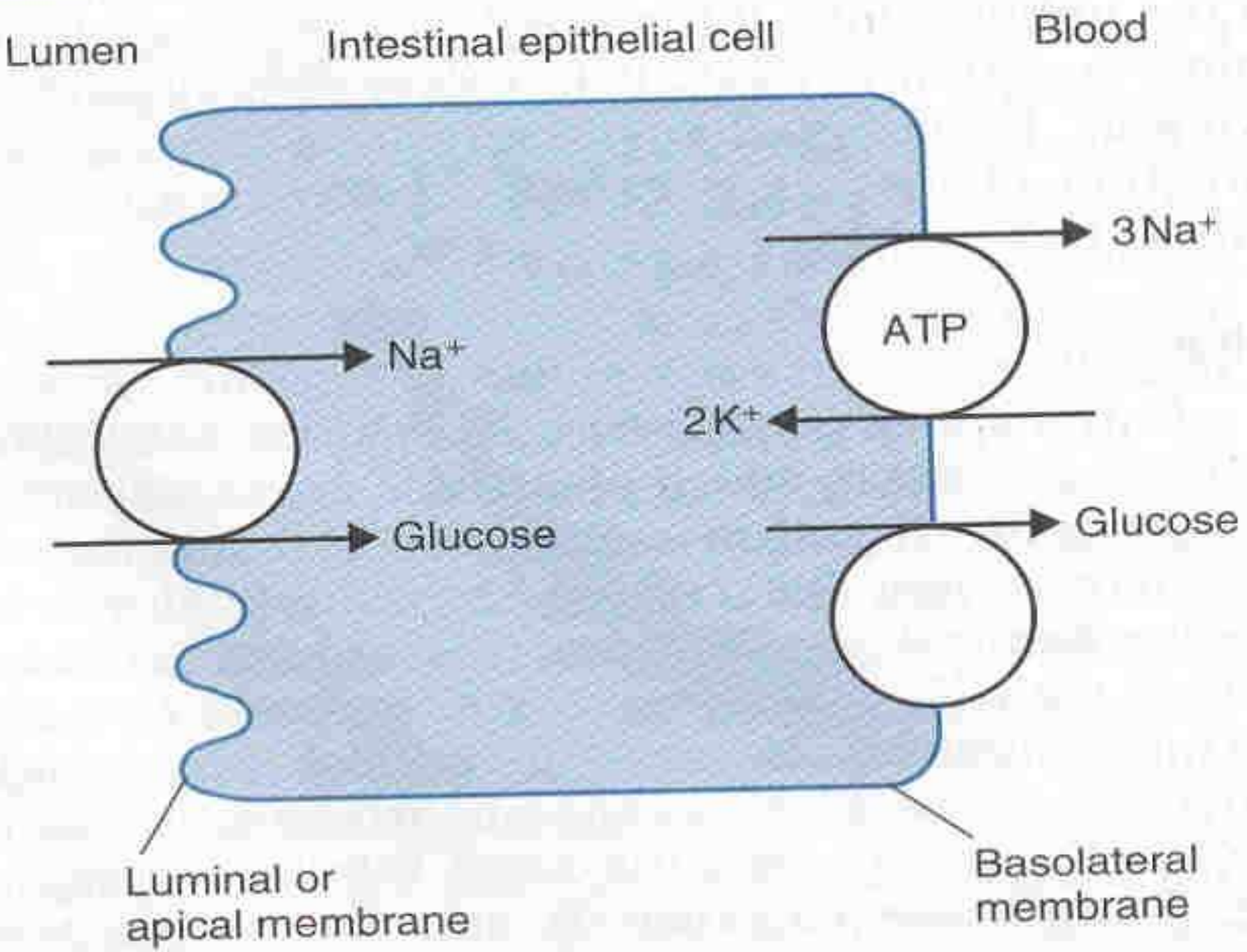
- **Co-transport:**

- All solutes move in the same direction ‘’ inside cell’’.

- e.g.
  - Na^+- glucose Co-transport.
  - Na^+- amino acid Co-transport.

  - present in the intestinal tract & kidneys.



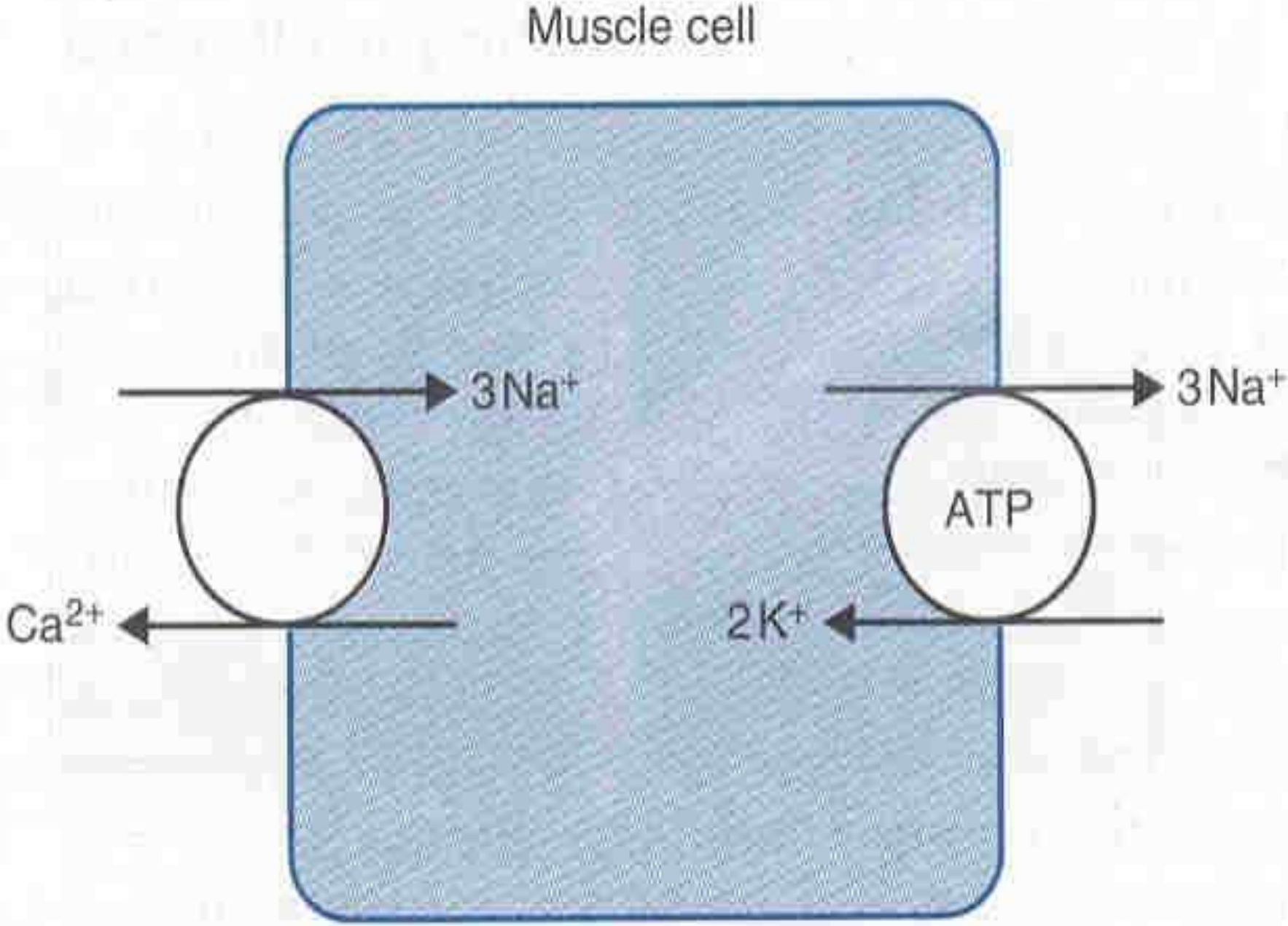

Lumen
Intestinal epithelial cell
Blood
3Na+
ATP
Na+
2K+
Glucose
Glucose
Luminal or
apical membrane
Basolateral
membrane

Counter-transport:

- Na^+ is moving to the interior of the cell causing other substance to move out.

- Ca^{++}- Na^+ exchange.
  (present in many cell membranes)

- Na^+- H^+ exchange in the kidney.


Muscle cell
3Na+
3Na+
ATP
Ca2+
2K+

Body Fluids & Electrolytes

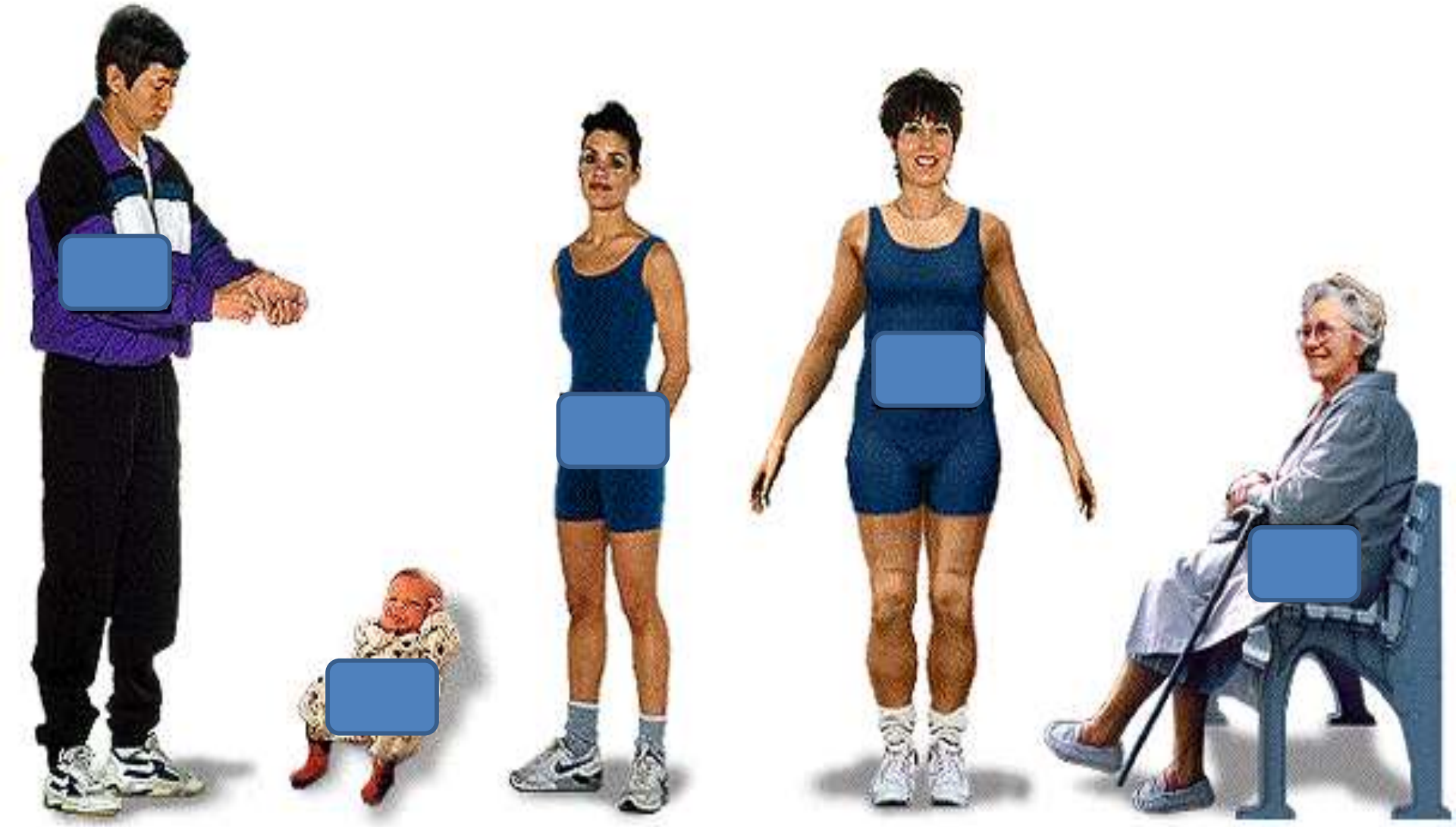
objectives

At the end of this session, the students should be able to:

- **Identify and describe daily intake and output of water and maintenance of water balance.**
- **List and describe of body fluid compartments as intra-cellular fluid (ICF) Extra-cellular fluid (ECF), interstitial fluid, trans-cellular fluid and total body water (TBW).**
- **Describe the composition of each fluid compartment, in terms of volume and ions and represent them in graphic forms.**
- **Physiology factor influencing body fluid: age, sex, adipose tissue, etc. Pathological factors: Dehydration, fluid infusion.**

- **Human body contain 50-70% water.**
- **E.g.**
  - **70 kg man has 42 L of water.**
  - **(Kg of water = Liter of water)**

Percentage of Water in the Body



FACTORS AFFECTING

Infant: 73%

Male adult: 60%

Female adult: 40-50%

Obesity

Old age 45%

Body Water Content

- **Infants** have low body fat, low bone mass, and are 73% or more water.
- Total water content declines throughout life.
- Healthy **males** are about **60% water**; healthy **females** are around **50%**
  - This difference reflects females':
    - Higher body fat
    - Smaller amount of skeletal muscle
- In **old age**, only about 45% of body weight is water.

Daily intake of water

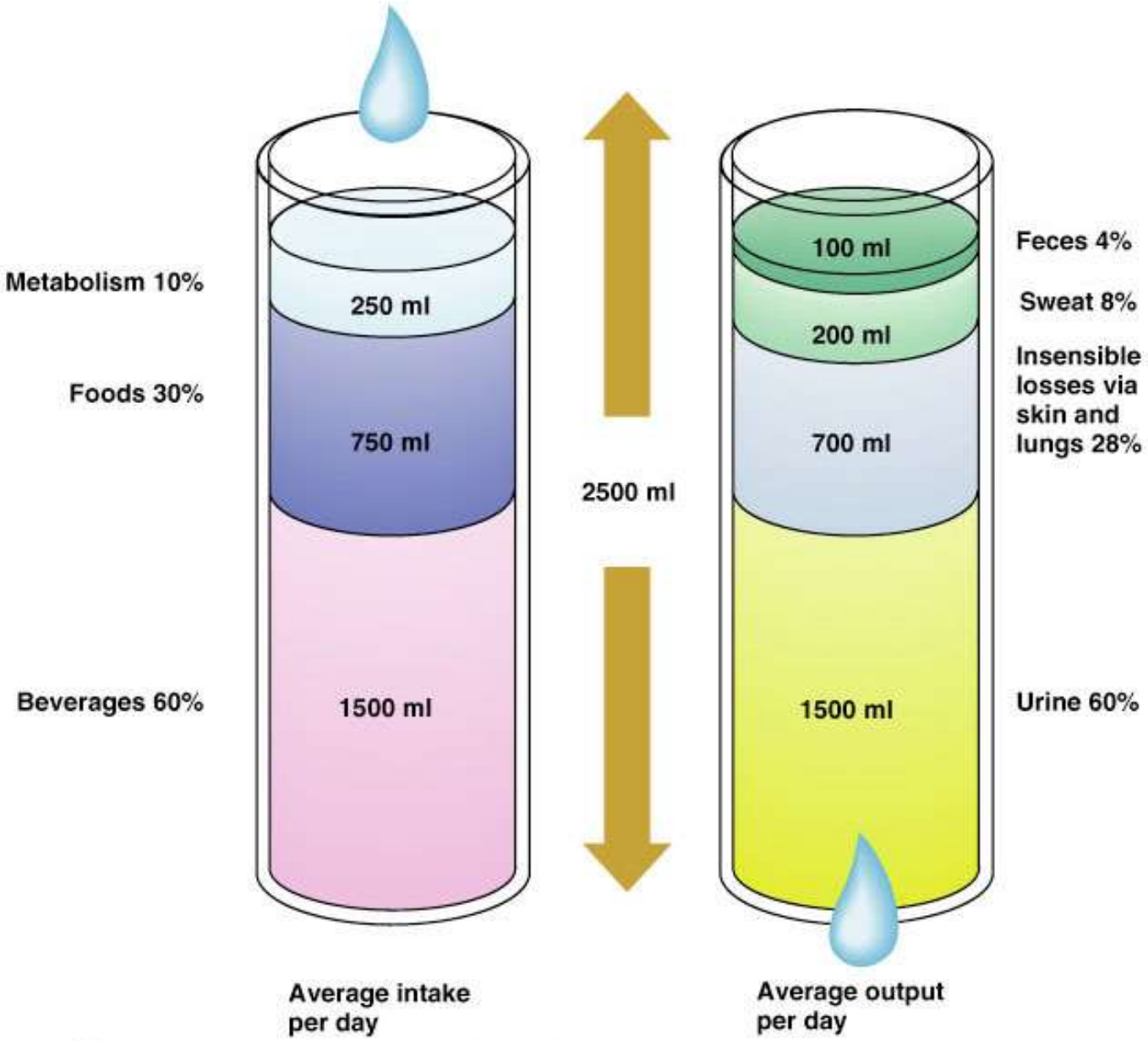
TABLE 20–1 DAILY INTAKE AND OUTPUT OF WATER (in ml/day)

	Normal	Prolonged, Heavy Exercise
Intake		
Fluids ingested	2100	?
From metabolism	200	200
Total intake	2300	?
Output		
Insensible—Skin	350	350
Insensible—Lungs	350	650
Sweat	100	5000
Feces	100	100
Urine	1400	500
Total output	2300	6600

Water Tank Analogy

Maintaining water homeostasis is a balancing act. The amount of water taken in must equal the amount of water lost.

Water Intake and Output



Regulation of Water Intake

Climate

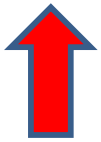
Habits

Level of physical activity.

Regulation of Water Intake

- The hypothalamic **thirst center** is stimulated:

– By a decline in plasma **volume** of 10%–15%



– By increases in plasma **osmolality** of 1–2%

In steady state water intake = water loss

Factors that affect the TBW

Physiological factors:

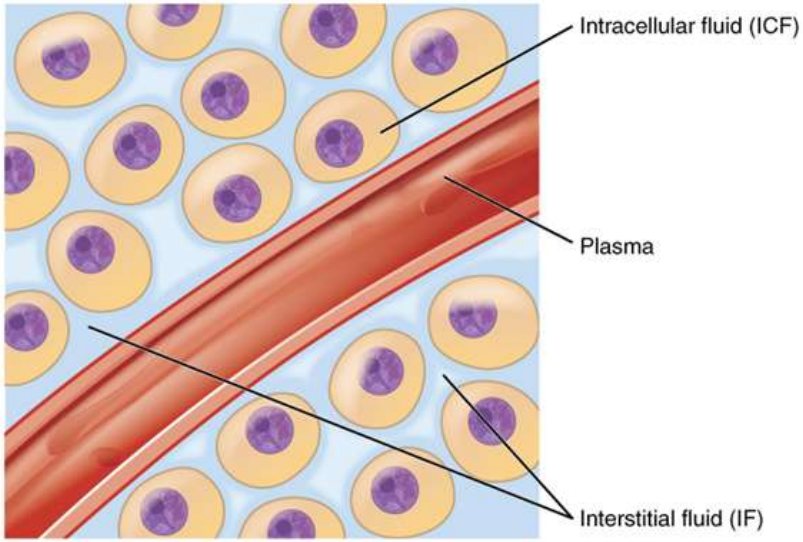
- Age
- Sex
- Body fat
- Climate
- Physical activity

Pathological factors:

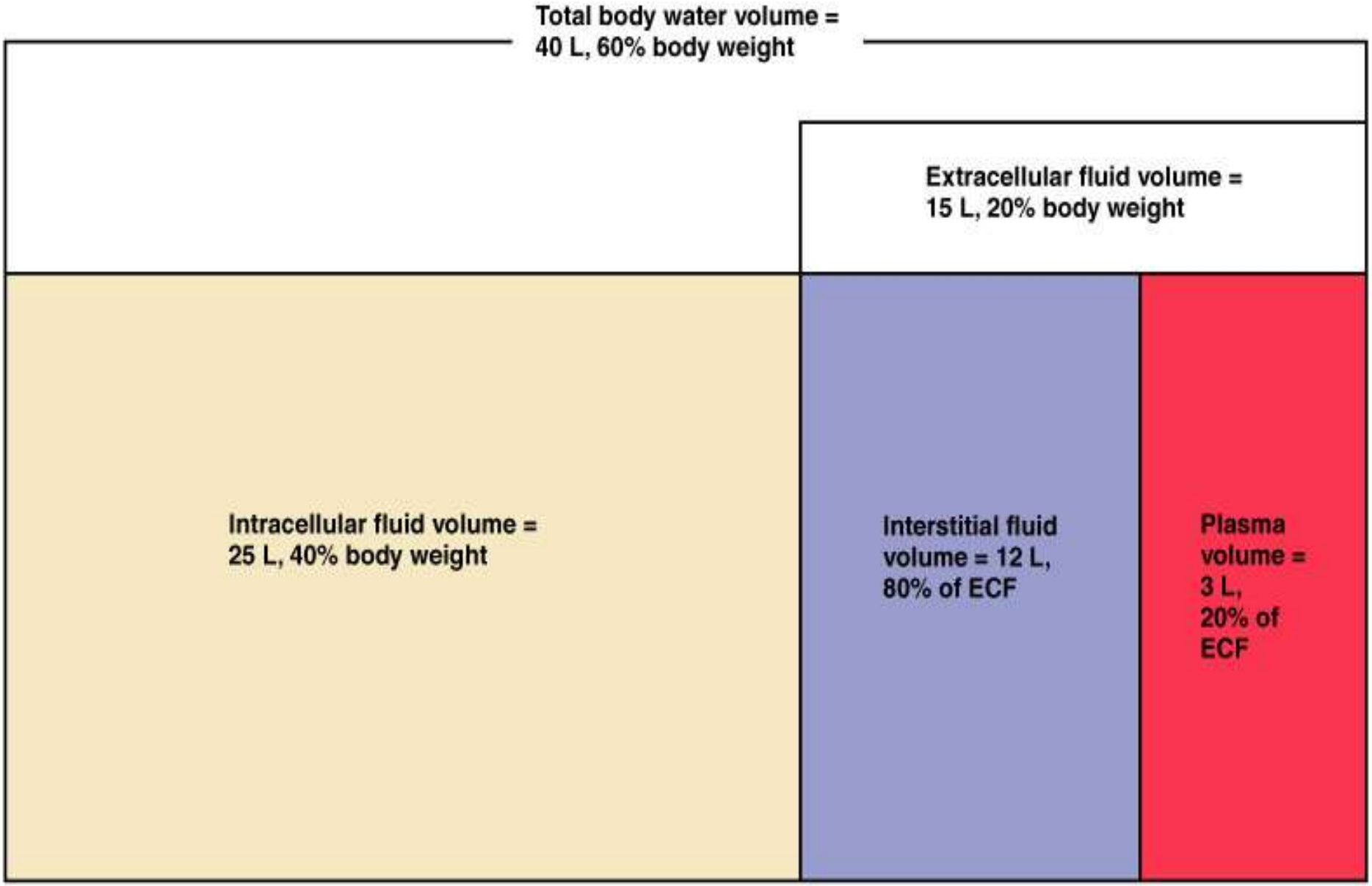
- Vomiting
- Diarrhea
- Diseases with excessive loss of water (DM, excessive sweating,….
- Blood loss

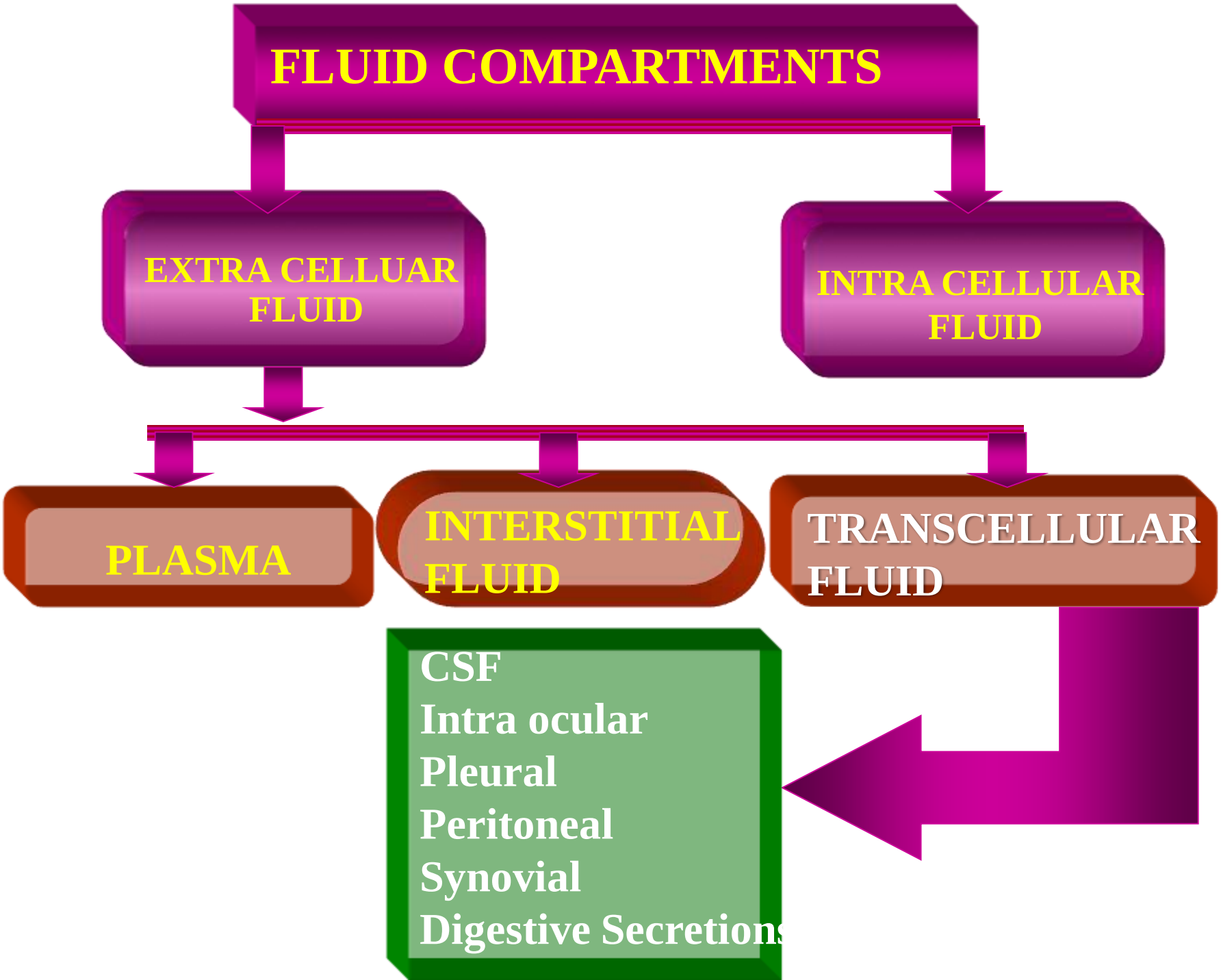
Fluid Compartments

- Water occupies two main fluid compartments:
  - Intracellular fluid (ICF)
  - Extracellular fluid (ECF)
    - Plasma
    - Interstitial fluid (IF)



Fluid Compartments




FLUID COMPARTMENTS
EXTRA CELLUAR FLUID
INTRA CELLULAR FLUID
PLASMA
INTERSTITIAL FLUID
TRANSCELLULAR FLUID
CSF
Intra ocular
Pleural
Peritoneal
Synovial
Digestive Secretions

Intracellular fluid (ICF)

- **Inside the cell.**
- **2/3 of TBW.**
- **High concentration of protein.**

Extracellular fluid (ECF)

- **Out side the cell.**
  1/3 of TBW.

1- Plasma:

Fluid circulating in the blood vessels.
1/4 of ECF

2- Interstitial fluid:

Fluid bathing the cell.
Ultra filtration of plasma.
3/4 of ECF

- **Plasma and interstitial fluid are almost having the same composition except for high protein concentration in plasma.**

Trancecellular fluid compartment:

- **small amount.**

CSF, GIT fluid, biliary fluid, synovial fluid, intrapelural fluid, intraperitoneal fluid, intrapericardial fluid and intraoccular fluid.

e.g.

- **TBW = 42L.**
- **ECF = 14L.**
- **ICF = 28L.**
- **Plasma = 3.5 L.**
- **Interstitial = 10.5 L.**

Composition of Body Fluids

- **Water** is the universal solvent.

- **Solutes** are broadly classified into:
  - **Electrolytes** – inorganic salts, all acids and bases, and some proteins
  - **Nonelectrolytes** – examples include glucose, lipids, creatinine, and urea

  - Amount = in **moles, osmoles.**

concentration

1- **Molarity** = moles/liter

(M/L)

2- **Osmolarity** = osmoles/liter

(osm/L)

3- **Osmolality** = osmoles/kg

(osm/kg)

In biological solutions:

- **Millimoles per liter (mM/L)**
- **Milliosmoles per (mOsm/L)**
- **1mM=1/1000 M**
- **1mOsm=1/1000 Osm**

Electrolyte Concentration

- Expressed in milliequivalents per liter (mEq/L), a measure of the number of electrical charges in one liter of solution.
- mEq/L = (concentration of ion in [mg/L]/the atomic weight of ion) × number of electrical charges on one ion.
- For single charged ions, 1 mEq = 1 mOsm
- For bivalent ions, 1 mEq = 1/2 mOsm

Constituents of ECF and ICF

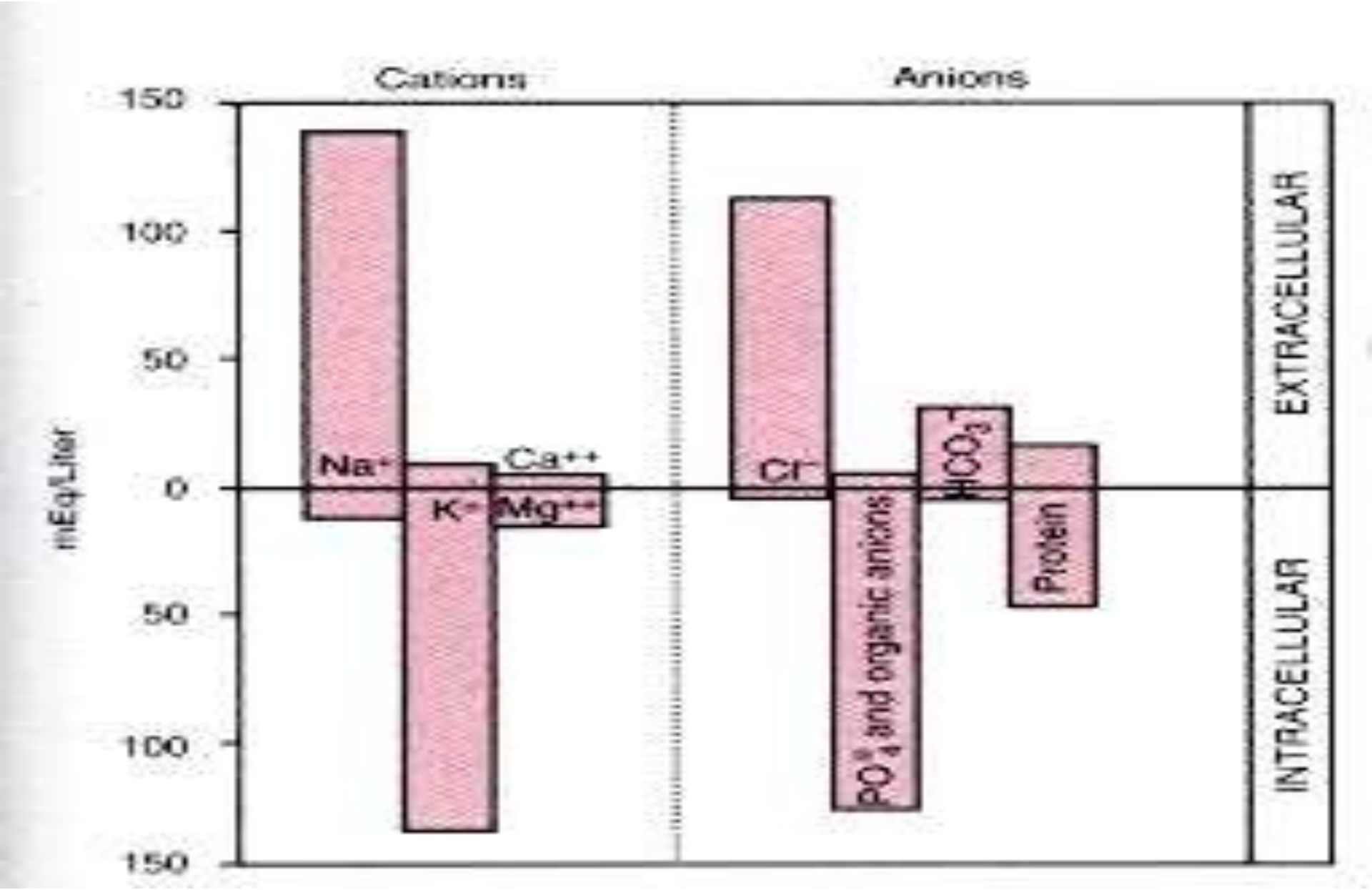


TABLE 20–2 OSMOLAR SUBSTANCES IN EXTRACELLULAR AND INTRACELLULAR FLUIDS

	Plasma (mOsm/liter of H_2O)	Interstitial	Intracellular
Na^+	142	139	14
K^+	4.2	4.0	140
Ca^{++}	1.3	1.2	0
Mg^{++}	0.8	0.7	20
Cl^-	108	108	4
HCO_3^-	24	28.3	10
HPO_4^{--}, $H_2PO_4^-$	2	2	11
SO_4^{--}	0.5	0.5	1
Phosphocreatine			45
Carnosine			14
Amino acids	2	2	8
Creatine	0.2	0.2	9
Lactate	1.2	1.2	1.5
Adenosine triphosphate			5
Hexose monophosphate			3.7
Glucose	5.6	5.6	
Protein	1.2	0.2	4
Urea	4	4	4
Others	4.8	3.9	10
Total mOsm/liter	301.8	300.8	301.2
Corrected osmolar activity (mOsm/liter)	282.0	281.0	281.0
Total osmotic pressure at 37° C (mm Hg)	5443	5423	5423

Extracellular and Intracellular Fluids

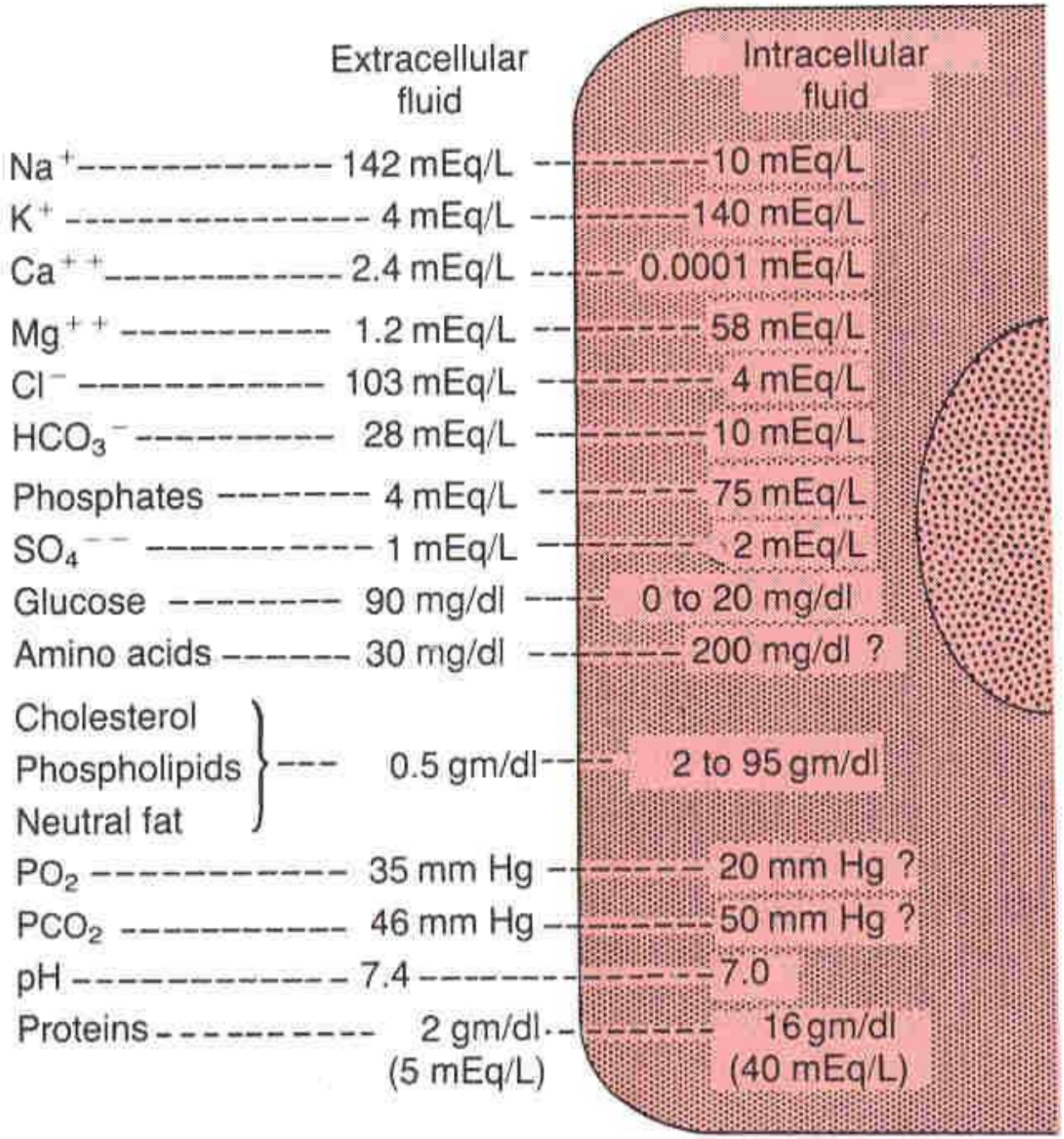
- Each fluid compartment of the body has a distinctive pattern of electrolytes.

- Extracellular fluids are similar (except for the high protein content of plasma)
  - Sodium is the chief cation
  - Chloride is the major anion

- **Intracellular fluid has low sodium and chloride**
  - **Potassium is the chief cation**
  - **Phosphate is the chief anion**

- **Each compartment must have almost the same concentration of positive charge (cations) as of negative charge (anion).**

(Electroneutrality)


Extracellular fluid
Intracellular fluid
Na^+ 142 mEq/L
10 mEq/L
K^+ 4 mEq/L
140 mEq/L
Ca^{++} 2.4 mEq/L
0.0001 mEq/L
Mg^{++} 1.2 mEq/L
58 mEq/L
Cl^- 103 mEq/L
4 mEq/L
HCO_3^- 28 mEq/L
10 mEq/L
Phosphates 4 mEq/L
75 mEq/L
SO_4^{--} 1 mEq/L
2 mEq/L
Glucose 90 mg/dl
0 to 20 mg/dl
Amino acids 30 mg/dl
200 mg/dl ?
Cholesterol
Phospholipids
Neutral fat
0.5 gm/dl
2 to 95 gm/dl
PO_2 35 mm Hg
20 mm Hg ?
PCO_2 46 mm Hg
50 mm Hg ?
pH 7.4
7.0
Proteins 2 gm/dl (5 mEq/L)
16 gm/dl (40 mEq/L)

- **Hypokalemia: decrease in K concentration in the ECF.**

1-2 mEq/L

- **Hyperkalemia: increase in K 60-100% above normal.**

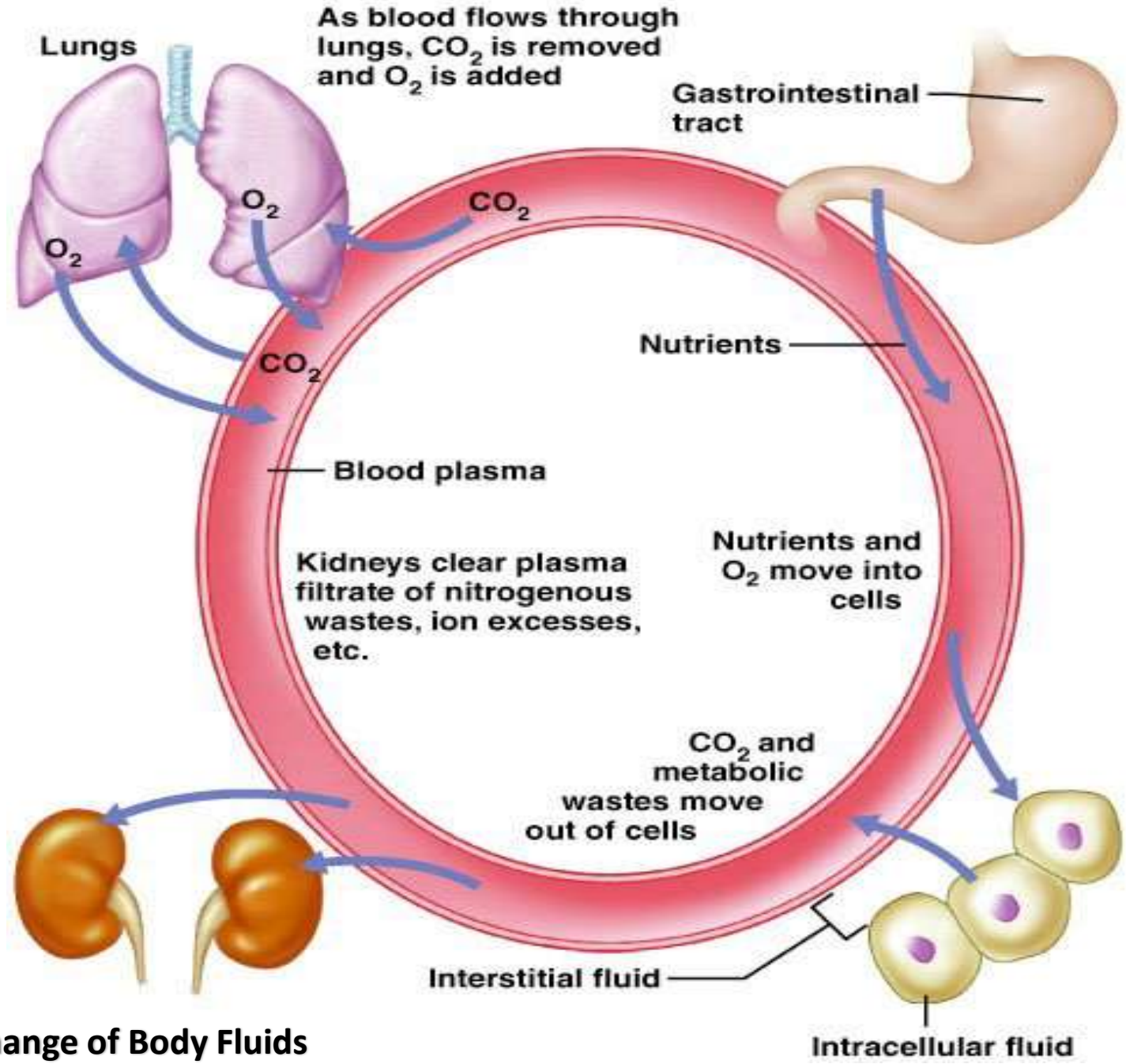
Hypernatremia:

increase **in Na concentration in ECF.**

Hyponatremia:

decrease **in Na concentration in the ECF.**

الله يوفقن



Continuous exchange of Body Fluids

Mechanisms for Movement

- **3 general mechanisms:**

1. **simple diffusion (passive)**
2. **Facilitated transport (passive)**
3. **Active transport**

osmosis

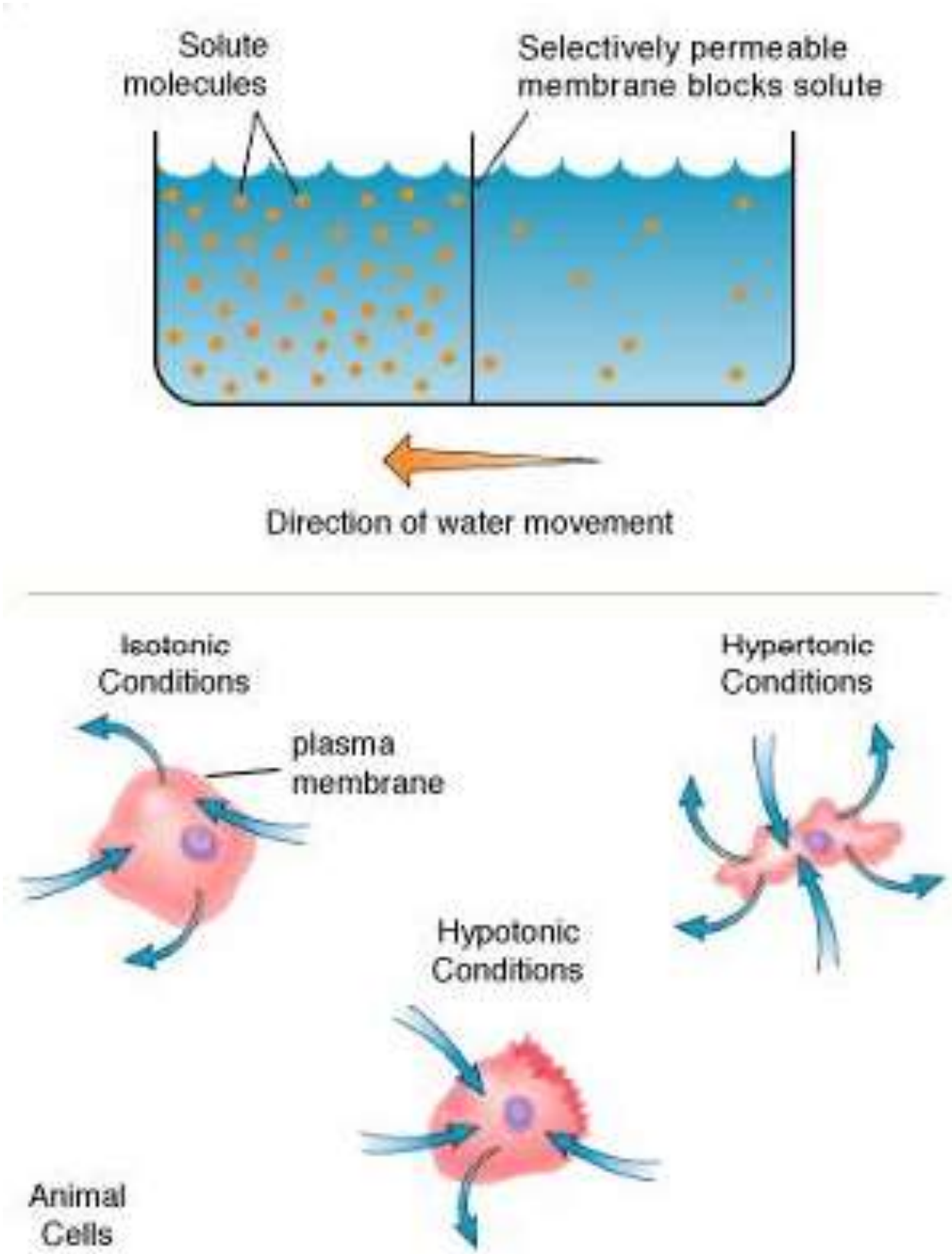
- **Net diffusion of water from a region of high water concentration to region of low water concentration.**

Osmotic equilibrium is maintained between intracellular and extracellular fluids:

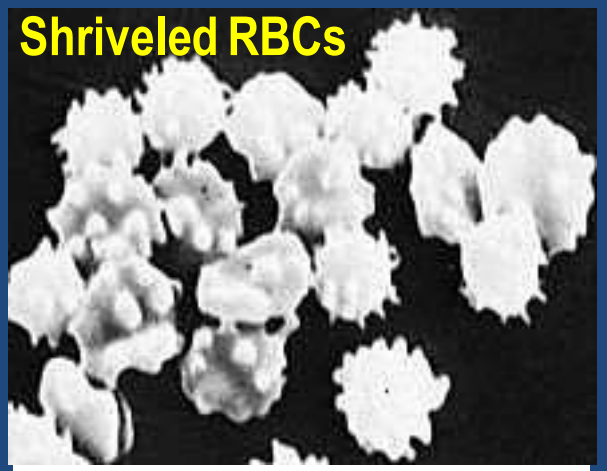
- **Small changes in concentration of solutes in the extracellular fluid can cause tremendous change in cell volume.**

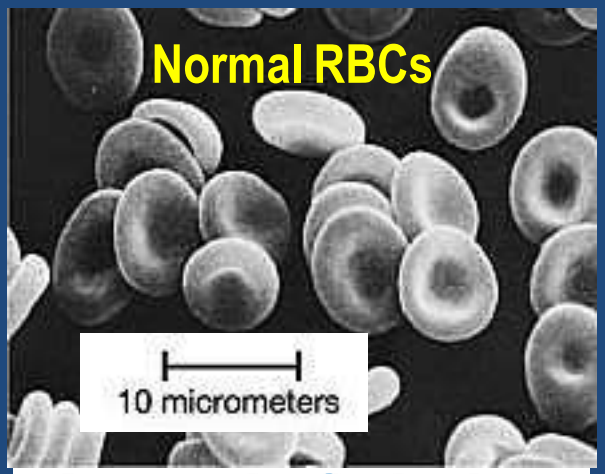
- **Intracellular osmolarity = extracellular osmolarity .**
- **≈ 300 mosm/L**

Osmosis

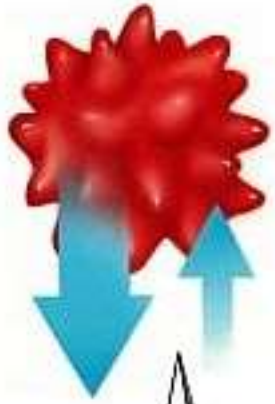

Solute molecules
Selectively permeable membrane blocks solute
Direction of water movement
Isotonic Conditions
plasma membrane
Hypertonic Conditions
Hypotonic Conditions
Animal Cells

Osmosis



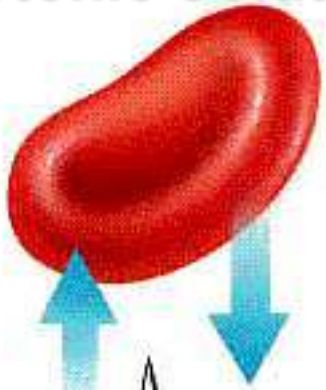


Hypertonic Solution



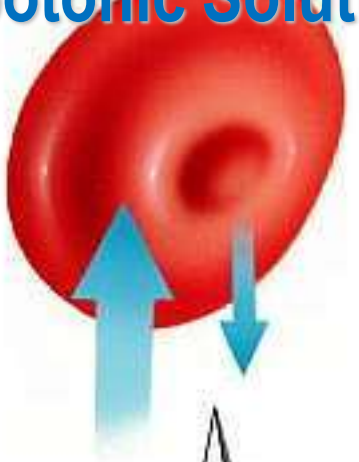
Net movement of water out of cells

Isotonic Solution



Equal movement of water into and out of cells

Hypotonic Solution



Net movement of water into cells

Osmosis

- **If environment is:**
  - **Hypertonic:**
    - **MORE SOLUTES outside cell**
    - **MORE WATER IN CELL**
    - **over time, cell loses water**
  - **Isotonic:**
    - **same**
    - **No change in cell volume**
  - **Hypotonic:**
    - **LESS SOLUTES outside cell**
    - **LESS WATER IN CELL, more solutes in cell.**
    - **over time, cell gains water**

❖ Isotonic solution :

- (not swell or shrink)
- 0.9% solution of sodium chloride or 5% glucose .
- same in and out .

❖ Hypotonic solution :

- (swelling) ⇓ 0.9%
- in is higher than out .

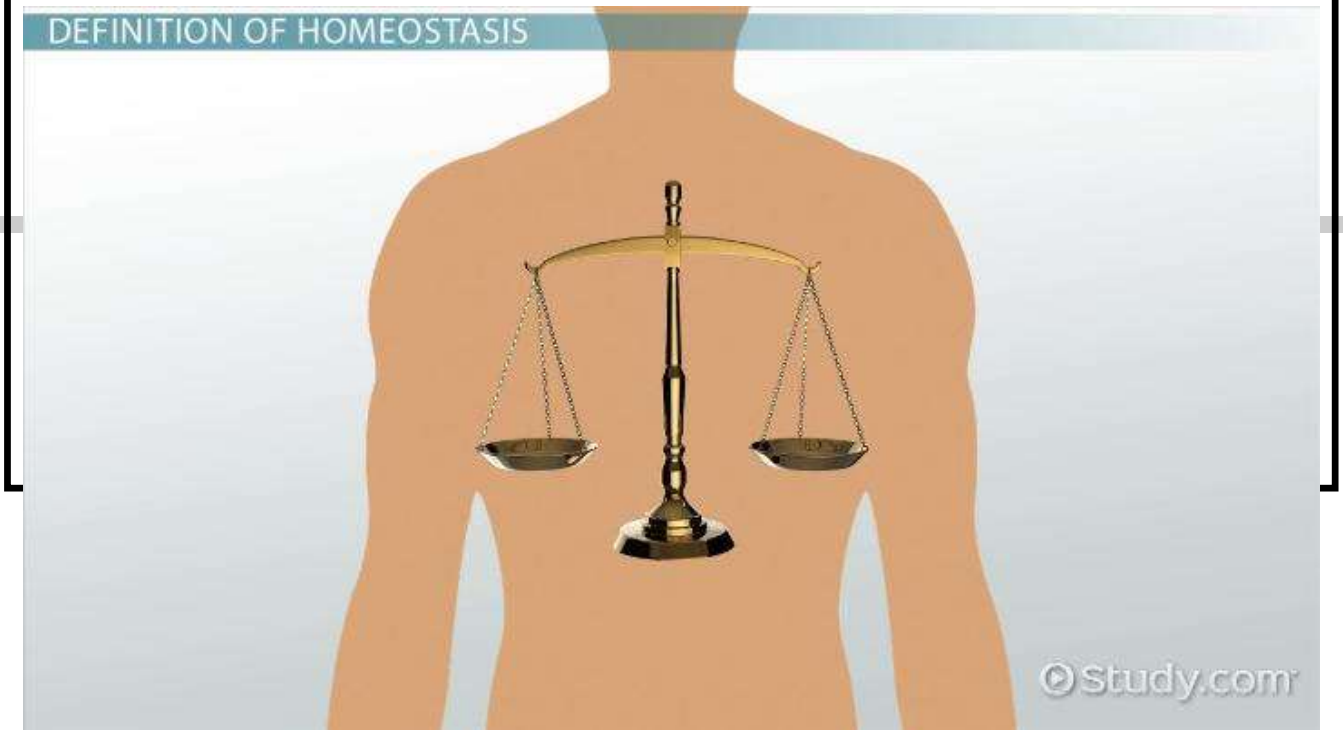
❖ Hypertonic solution :

- (shrink) ⇑ 0.9%
- out is higher than in

Glucose and other solutions administered for nutritive purposes

- **People who can not take adequate amount of food.**
- **Slowly.**
- **Prepared in isotonic solution.**

Homeostasis



Homeostasis

- **Homeostasis is the ability to maintain a relatively stable internal environment in an ever-changing outside world**
- **The internal environment of the body (ECF)is in a dynamic state of equilibrium**
- **All different body systems operate in harmony to provide homeostasis**

Homeostatic Control Mechanisms

- **The variable produces a change in the body**

- **The three interdependent components of control mechanisms are:**
  - **Receptor – monitors the environments and responds to changes (stimuli)**
  - **Control center – determines the set point at which the variable is maintained**
  - **Effector – provides the means to respond to the stimulus**

Regulation of body functions

1. Nervous system

- sensory input.
- central nervous system.
- motor out put.

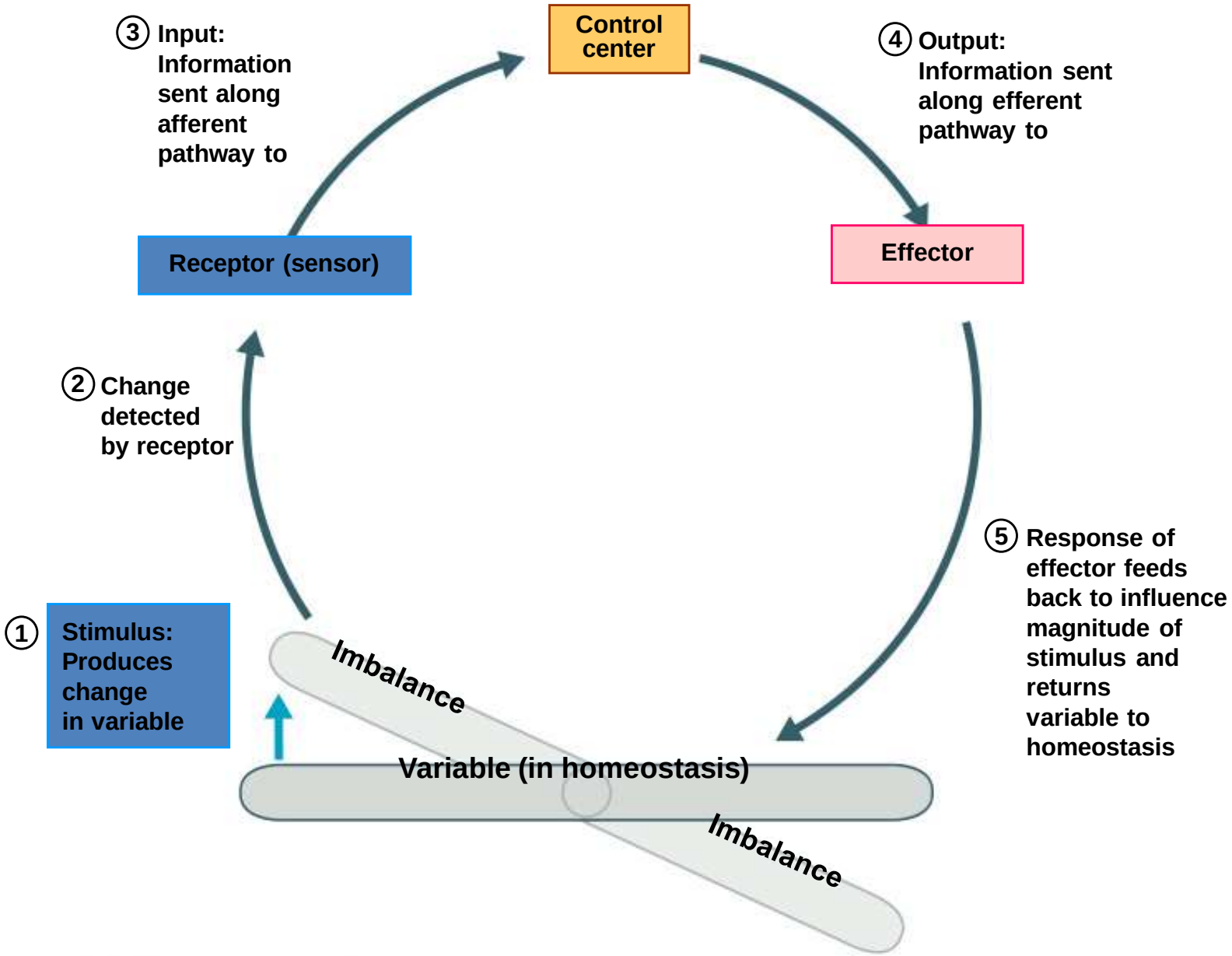
2. Hormonal system of regulation.

- **Endocrine gland.**

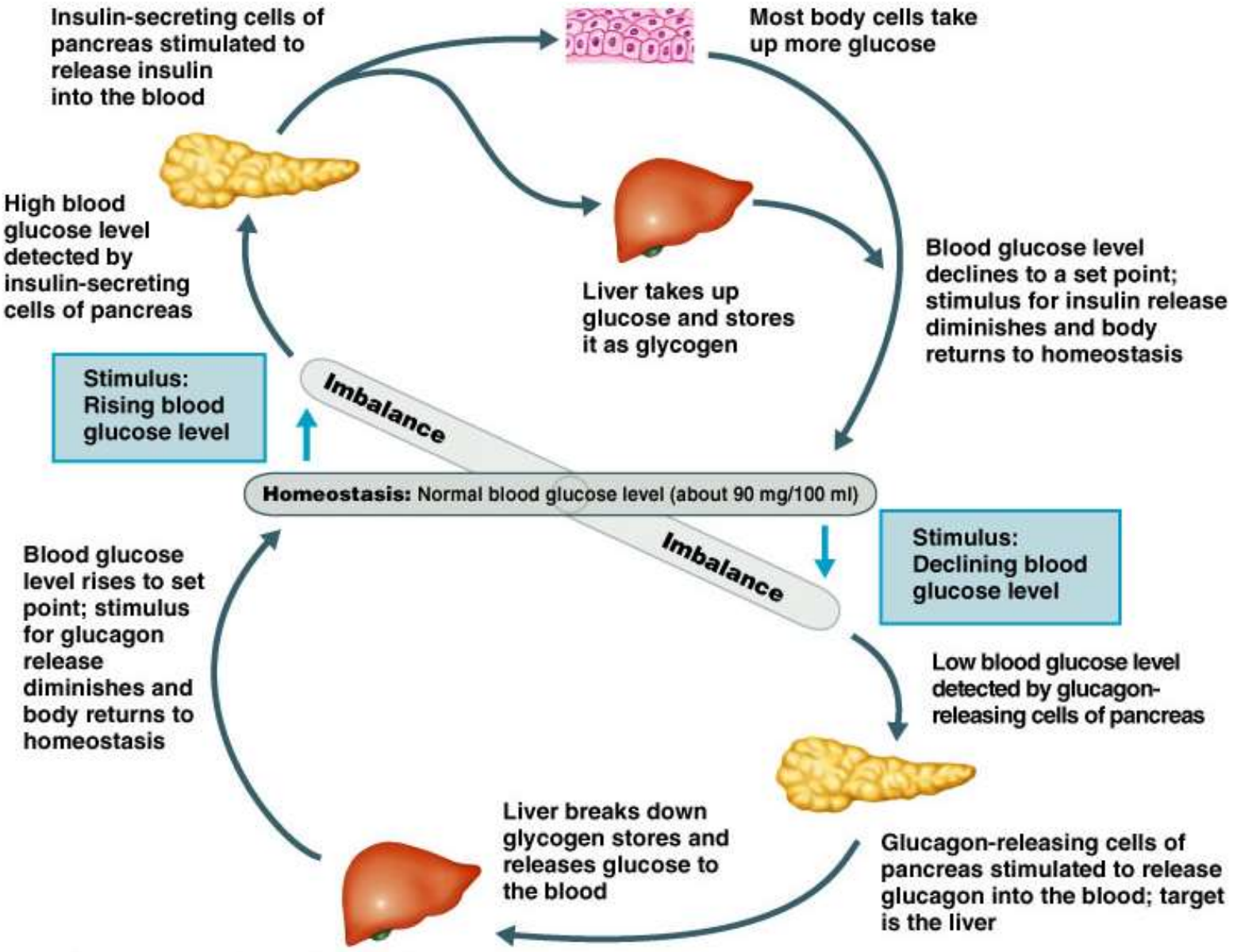
  Pancreas, thyroid

e.g. : insulin control glucose level.

Homeostatic Control Mechanisms



Feedback

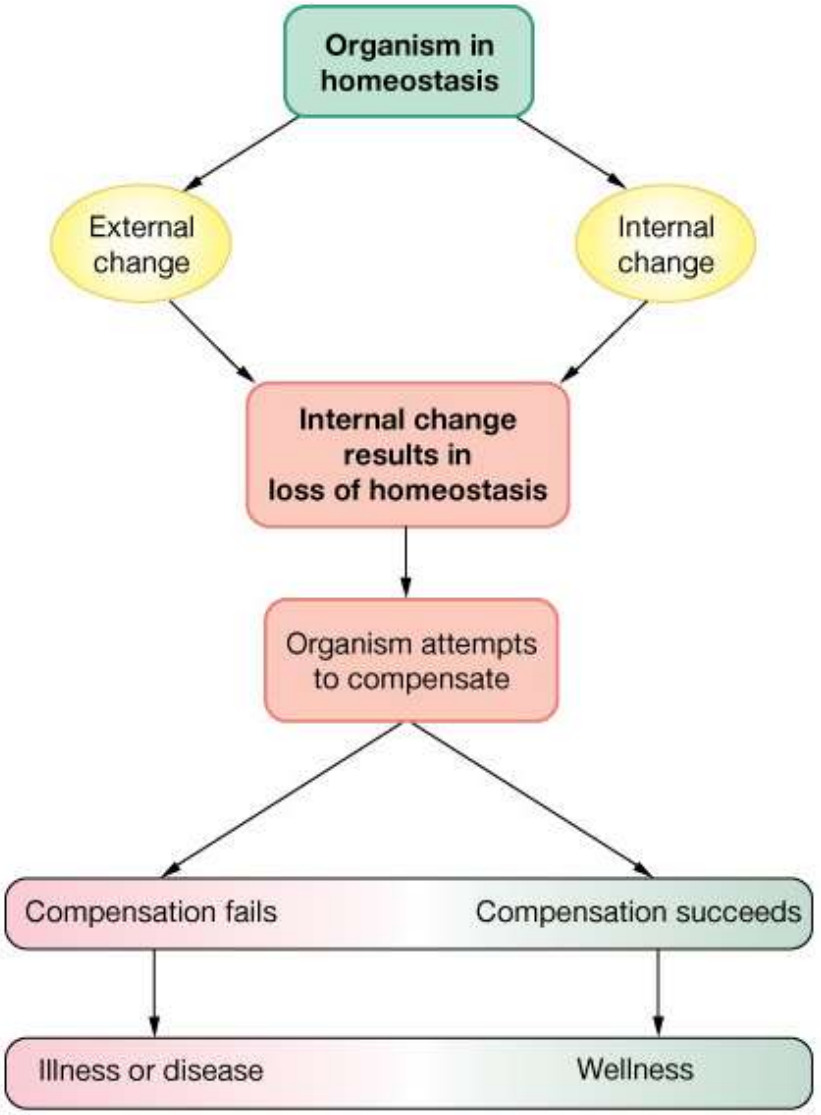

Insulin-secreting cells of pancreas stimulated to release insulin into the blood
Most body cells take up more glucose
High blood glucose level detected by insulin-secreting cells of pancreas
Liver takes up glucose and stores it as glycogen
Blood glucose level declines to a set point; stimulus for insulin release diminishes and body returns to homeostasis
Stimulus: Rising blood glucose level
Imbalance
Homeostasis: Normal blood glucose level (about 90 mg/100 ml)
Imbalance
Stimulus: Declining blood glucose level
Blood glucose level rises to set point; stimulus for glucagon release diminishes and body returns to homeostasis
Low blood glucose level detected by glucagon-releasing cells of pancreas
Liver breaks down glycogen stores and releases glucose to the blood
Glucagon-releasing cells of pancreas stimulated to release glucagon into the blood; target is the liver

Homeostatic Imbalance

- **Disturbance of homeostasis or the body's normal equilibrium.**

Homeostasis & Controls

- **Successful compensation**
  - Homeostasis reestablished
- **Failure to compensate**
  - Pathophysiology
    - Illness
    - Death



Lecture 4

Changes in The Body Fluid Compartments (ECF & ICF) and Edema

Fluid Compartments

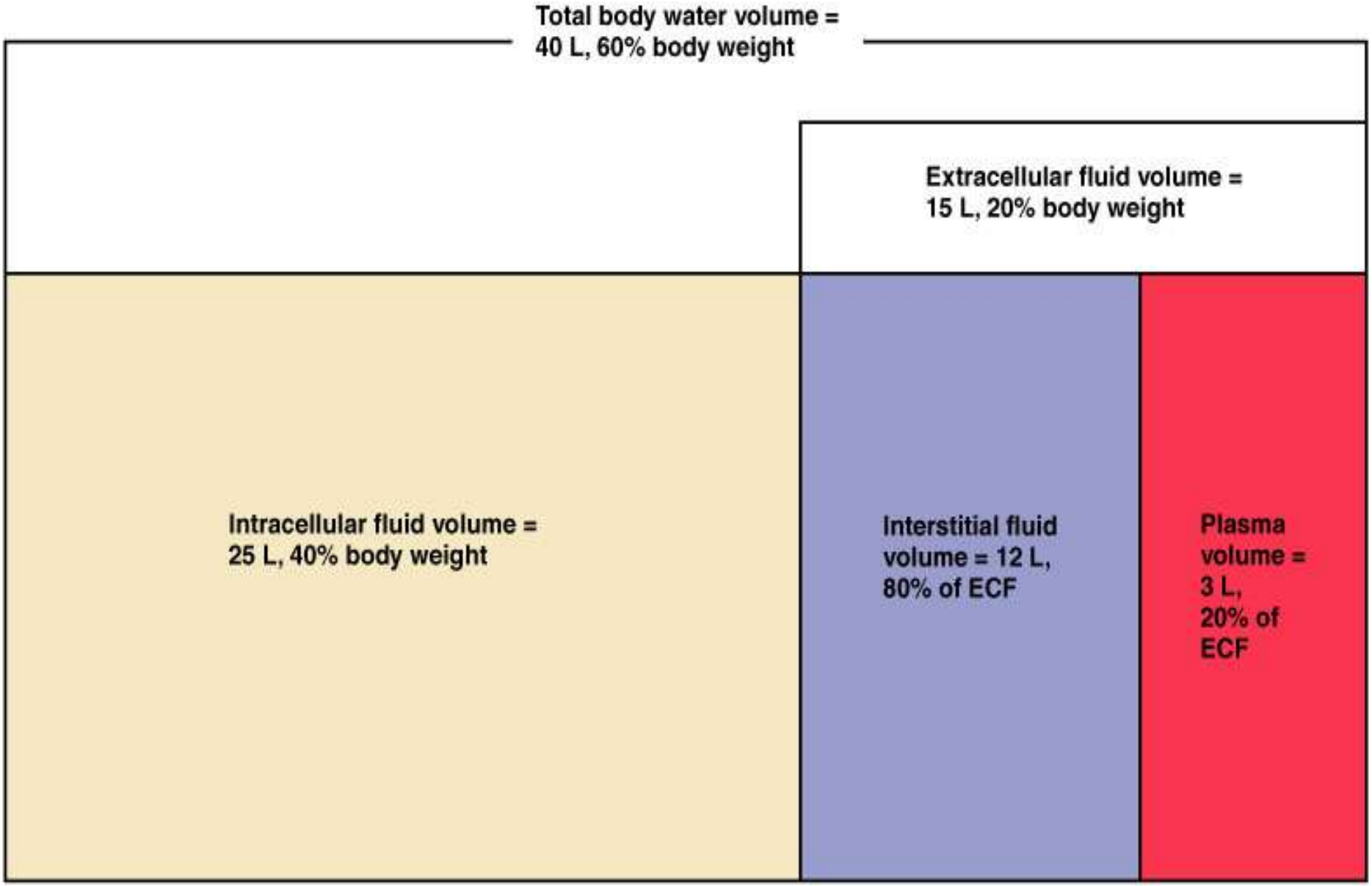
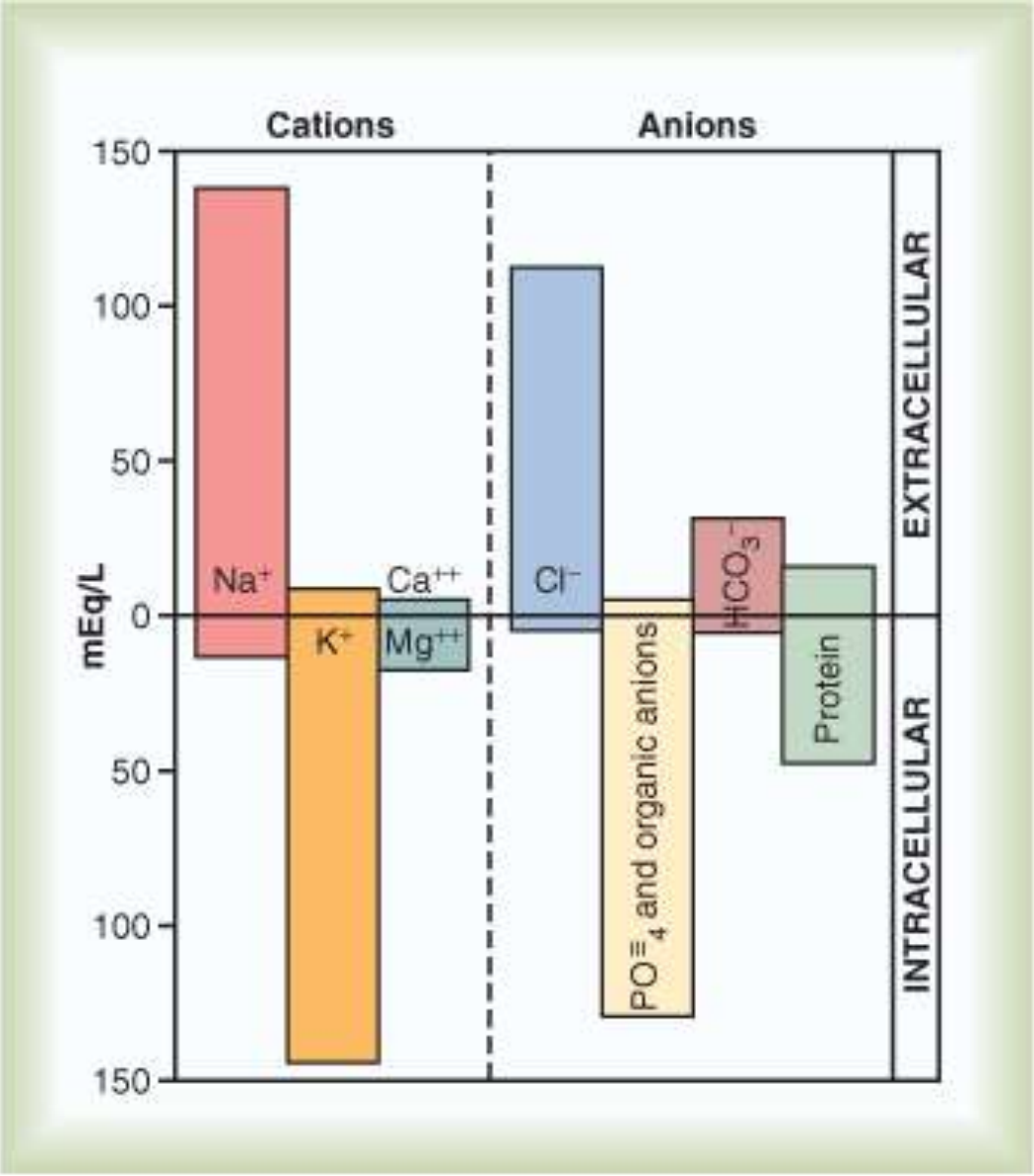


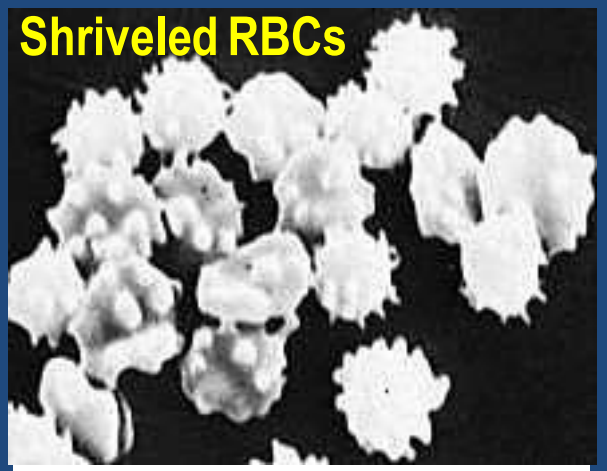
TABLE 20–2 OSMOLAR SUBSTANCES IN EXTRACELLULAR AND INTRACELLULAR FLUIDS

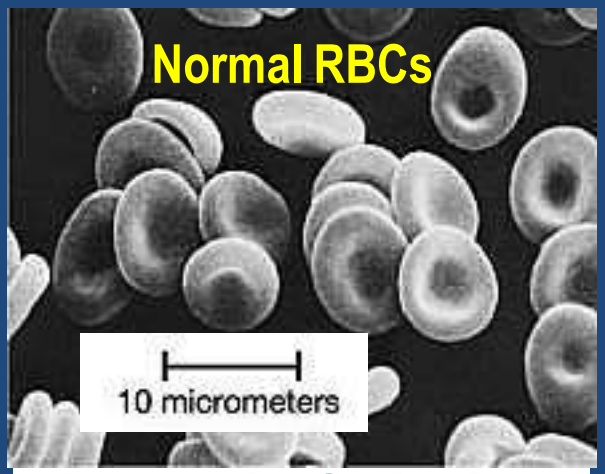
	Plasma (mOsm/liter of H_2O)	Interstitial	Intracellular
Na^+	142	139	14
K^+	4.2	4.0	140
Ca^{++}	1.3	1.2	0
Mg^{++}	0.8	0.7	20
Cl^-	108	108	4
HCO_3^-	24	28.3	10
HPO_4^{--}, $H_2PO_4^-$	2	2	11
SO_4^{--}	0.5	0.5	1
Phosphocreatine			45
Carnosine			14
Amino acids	2	2	8
Creatine	0.2	0.2	9
Lactate	1.2	1.2	1.5
Adenosine triphosphate			5
Hexose monophosphate			3.7
Glucose	5.6	5.6	
Protein	1.2	0.2	4
Urea	4	4	4
Others	4.8	3.9	10
Total mOsm/liter	301.8	300.8	301.2
Corrected osmolar activity (mOsm/liter)	282.0	281.0	281.0
Total osmotic pressure at 37° C (mm Hg)	5443	5423	5423

Constituents of ECF and ICF

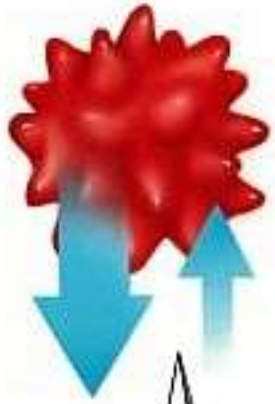


Osmosis



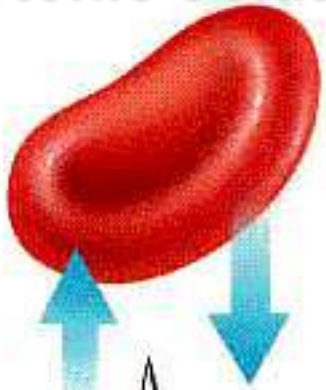


Hypertonic Solution



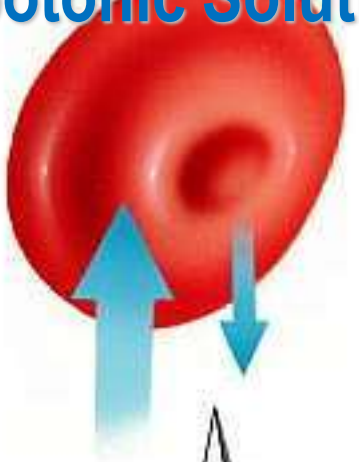
Net movement of water out of cells

Isotonic Solution



Equal movement of water into and out of cells

Hypotonic Solution



Net movement of water into cells

Volumes And Osmolarities of ECF and ICF In Abnormal States.

- **Some factors can cause the change:**
  - dehydration
  - intravenous infusion (IV)
  - abnormal sweating.
  - etc..

- **Changes in volume :**

1. **Volume contraction.**

2. **Volume expansion.**

Changes in volume

Volume contraction

removing

1- ***isotonic*** solution.

2- ***hypertonic*** solution.

3- ***hypotonic*** solution.

Volume expansion

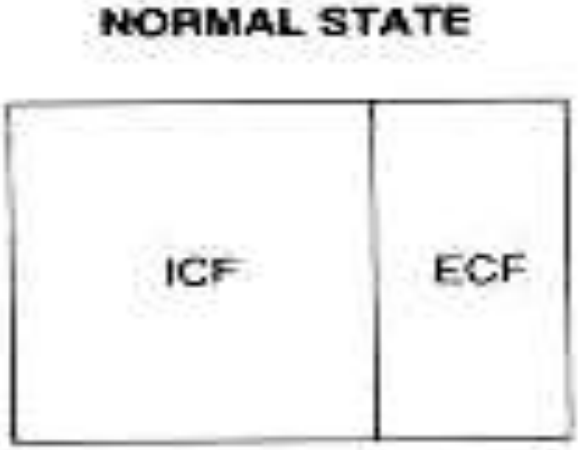
Adding

1- ***isotonic*** solution.

2- ***hypertonic*** solution.

3- ***hypotonic*** solution.

1- Loss of iso-osmotic fluid e.g. Diarrhea



Volume contraction:

1. **Diarrhea.**

   - **osmolarity of fluid lost ≈ osmolarity of ECF**

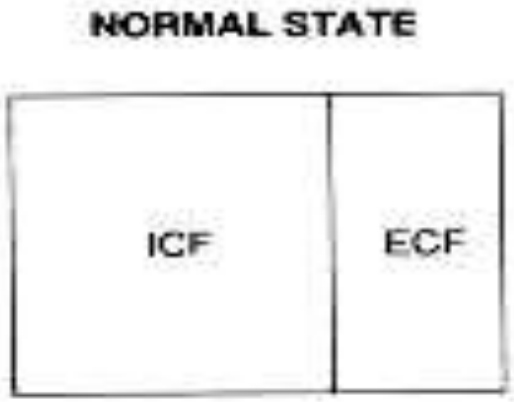
   (loss of isosmotic fluid).

   - ↓ **volume in ECF.**

   - ↓ **arterial pressure.**

2. Loss of hypotonic solution

e.g. Water deprivation



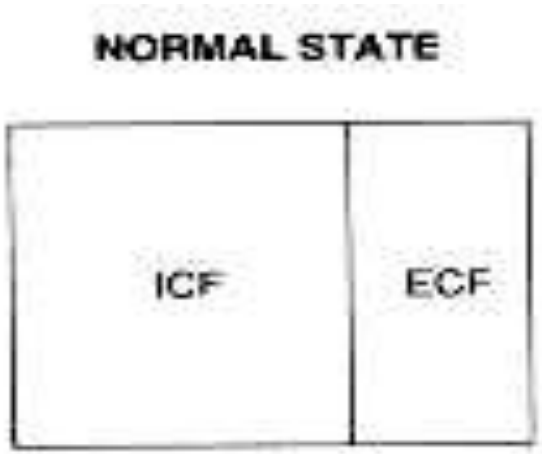
2. Water deprivation :

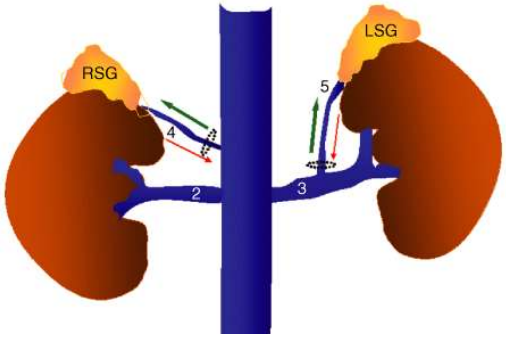
- Osmolarity and volume will change .

- ↑ Osmolarity in both ECF and ICF.

- ↓ Volume in both ECF and ICF.

3- Loss of hypertonic sol. e.g. Adrenal insufficiency





3. Loss of hypertonic solution

e.g. Adrenal insufficiency:

i.e. Aldosterone deficiency.

↓- Na^+ in the ECF.

↓- osmolarity in both .

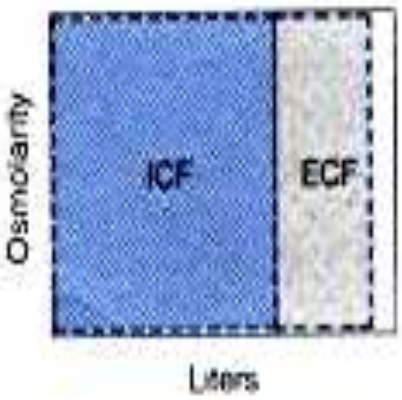
↓- in ECF volume.

↑- in ICF volume.

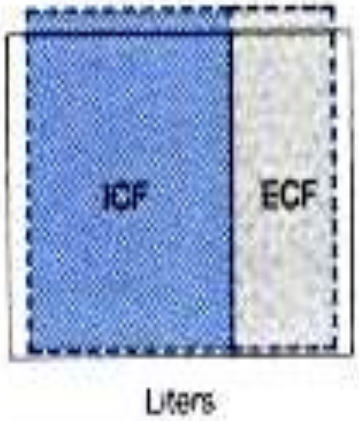
NORMAL STATE

VOLUME CONTRACTION

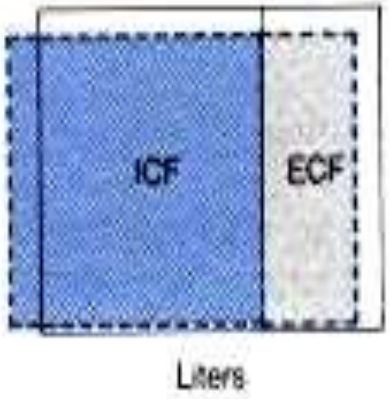
Diarrhea



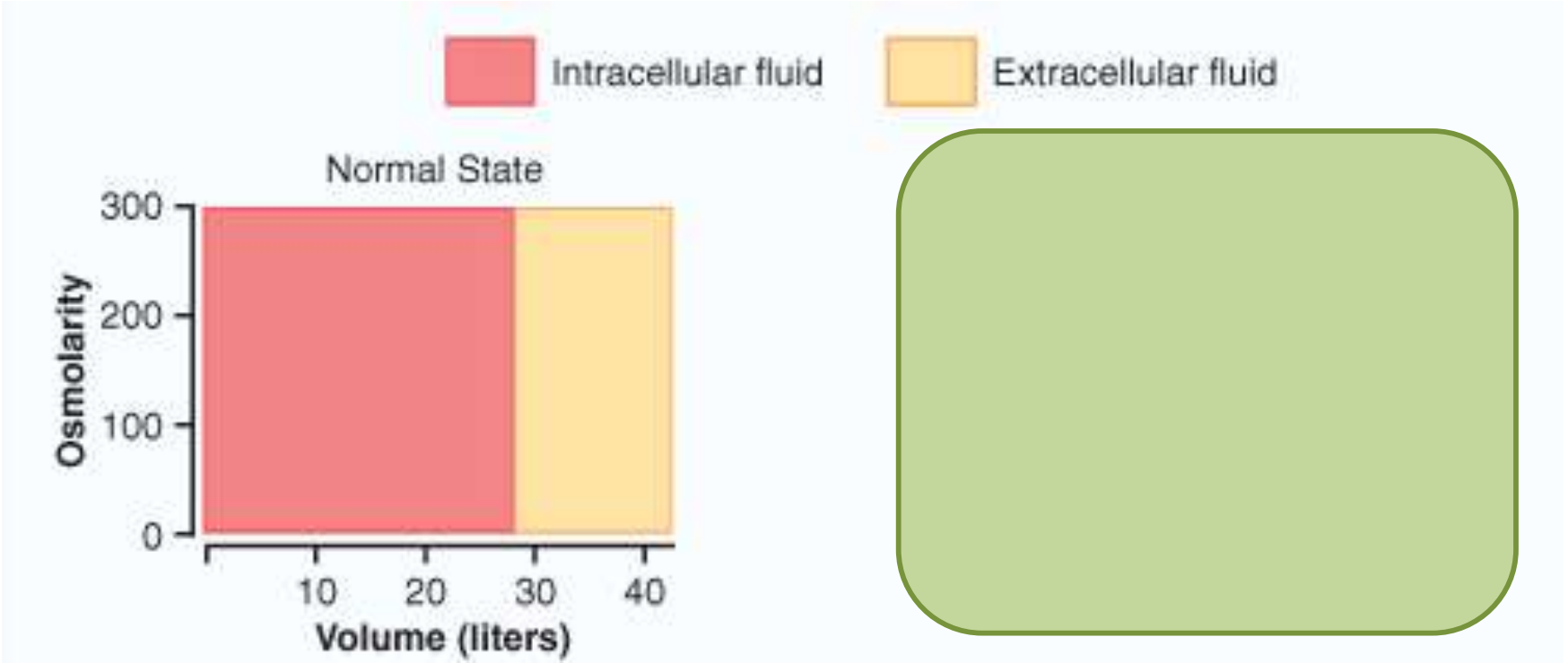
Water deprivation



Adrenal insufficiency



Volume Expansion



1. **Adding of isotonic NaCl.**

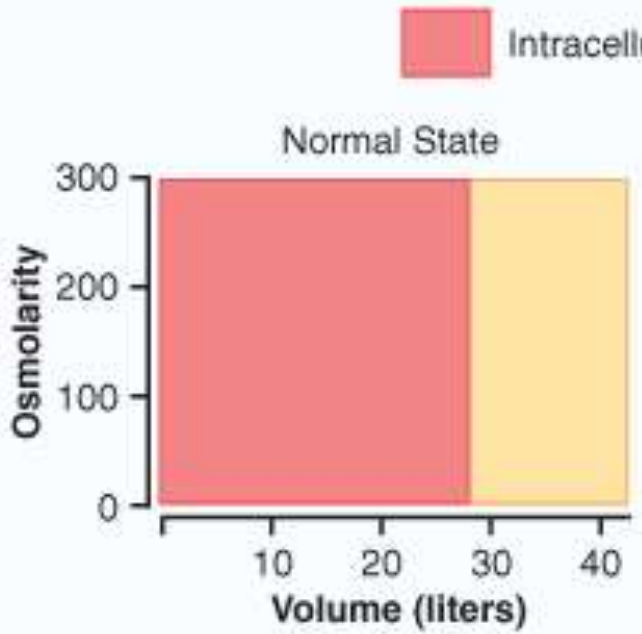
Volume Expansion

1. **Infusion of isotonic NaCl.**

↑ - in ECF volume.

- No change in osmolarity.

- ***Isomotic expansion*** .



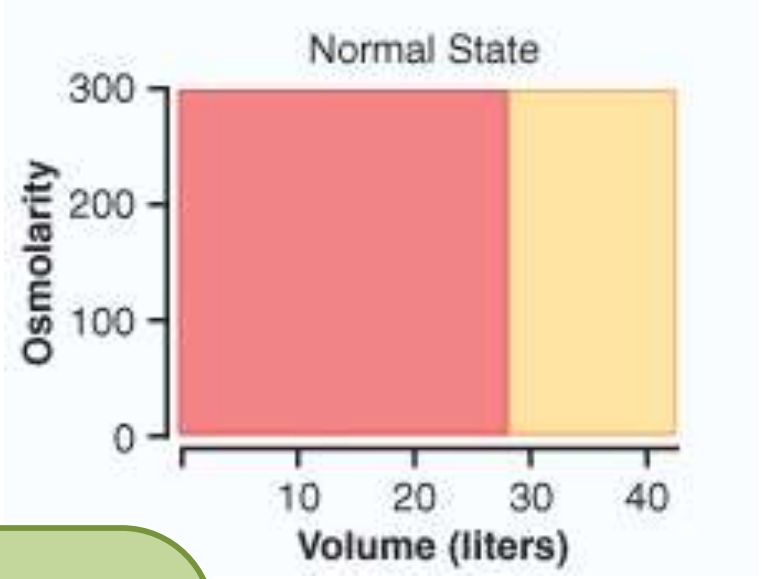
2- High NaCl intake

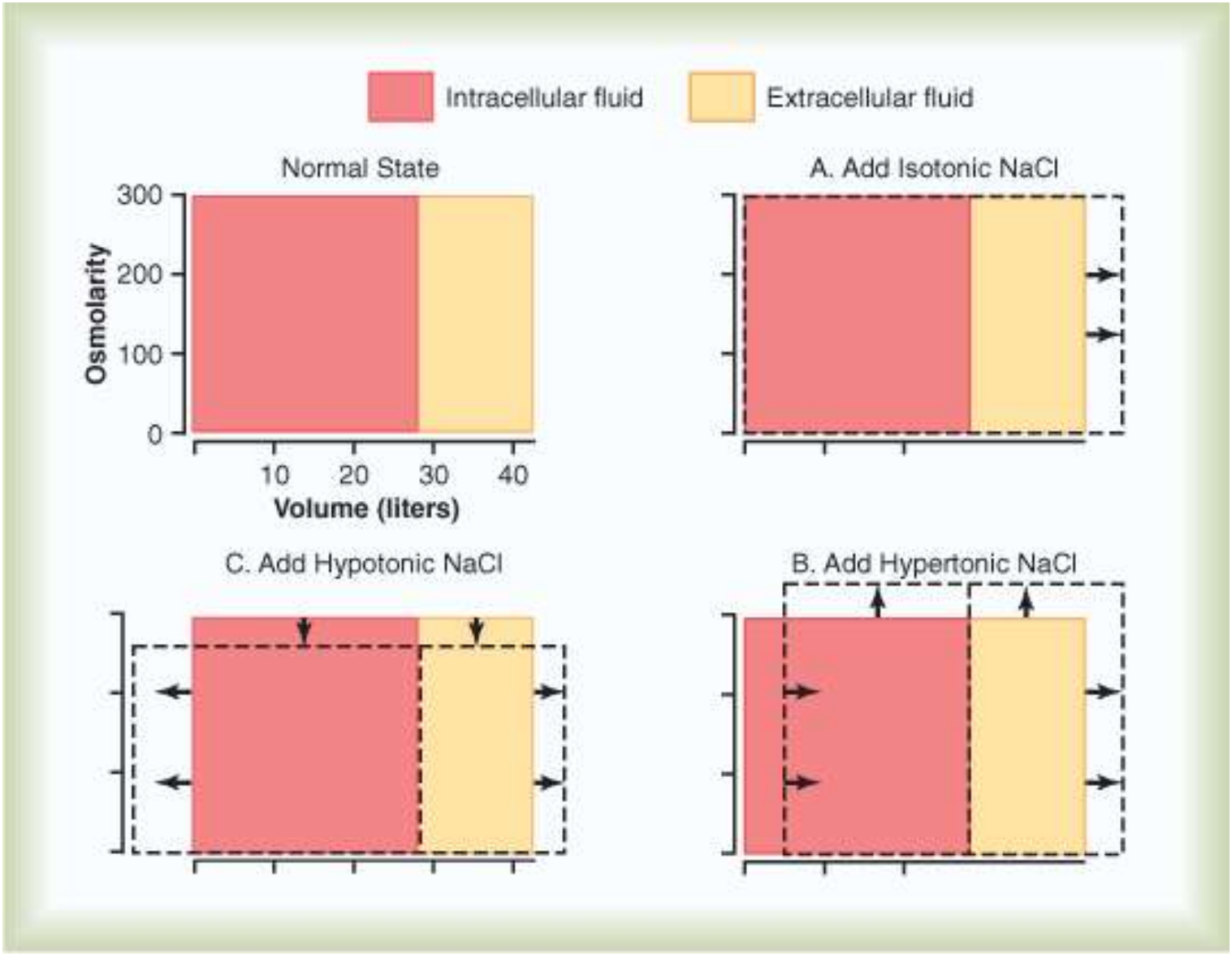
2. High NaCl intake.

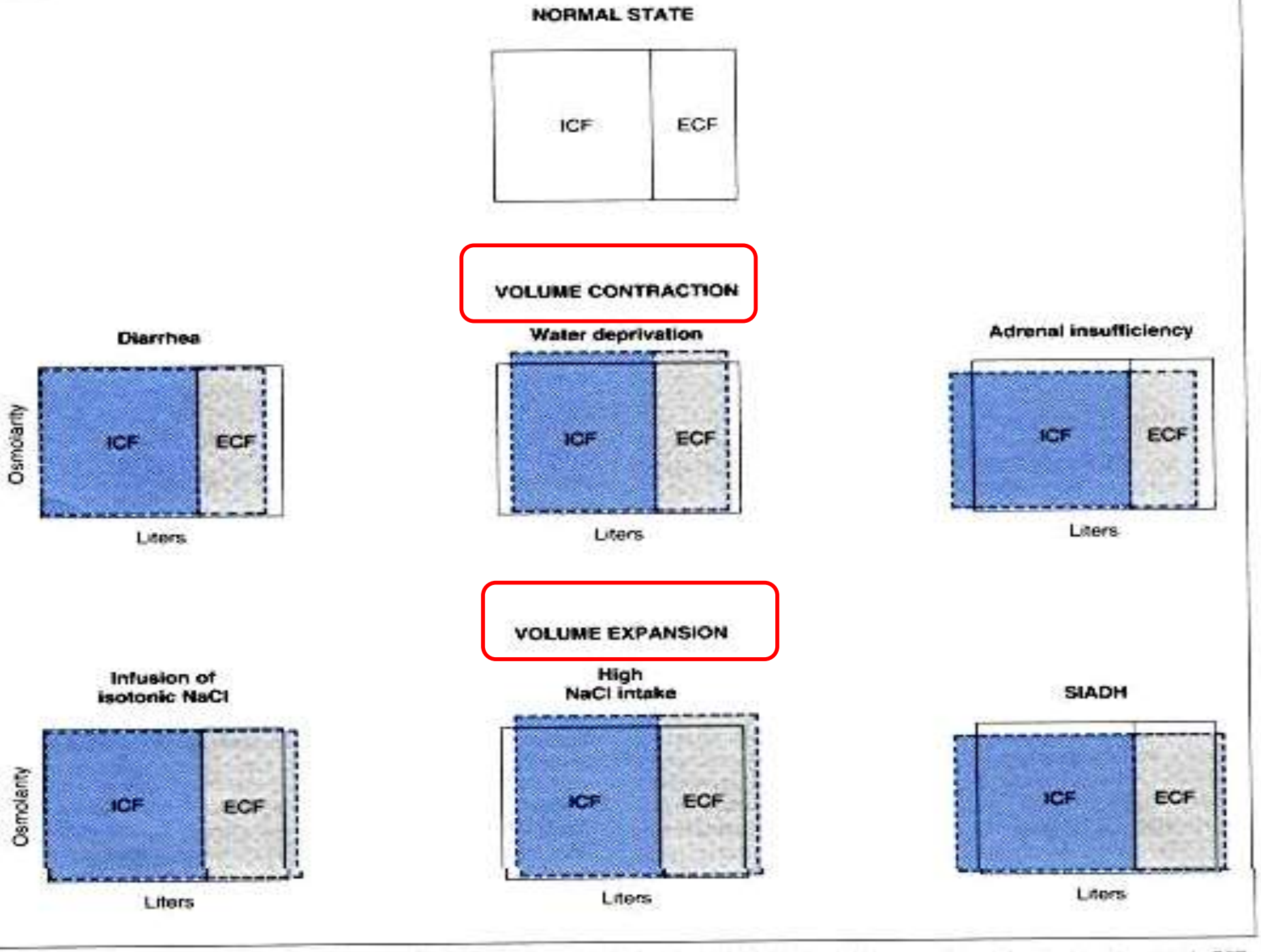
- ↑ eating salt.
- ↑ osmolarity in both.
- ↓ volume of ICF .
- ↑ volume of ECF .
- ***hyperosmotic volume expansion***.

3- **Adding hypotonic solution e.g. Syndrome of inappropriate antidiurtic hormone (SIADH)**

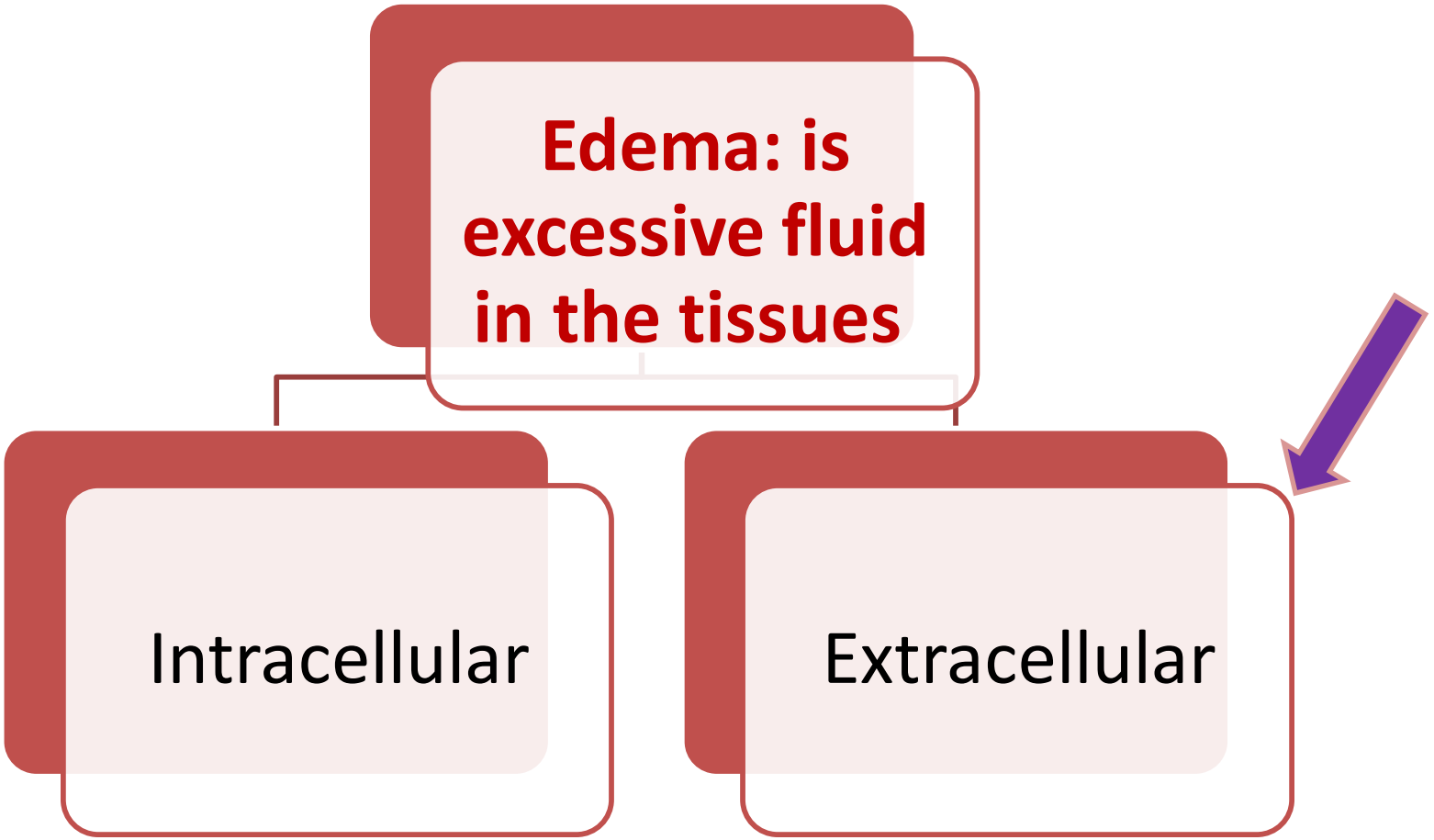
- ↑ volume
- ↓ osmolarity






NORMAL STATE
ICF
ECF
VOLUME CONTRACTION
Diarrhea
Water deprivation
Adrenal insufficiency
Osmolarity
ICF
ECF
Liters
VOLUME EXPANSION
Infusion of isotonic NaCl
High NaCl intake
SIADH

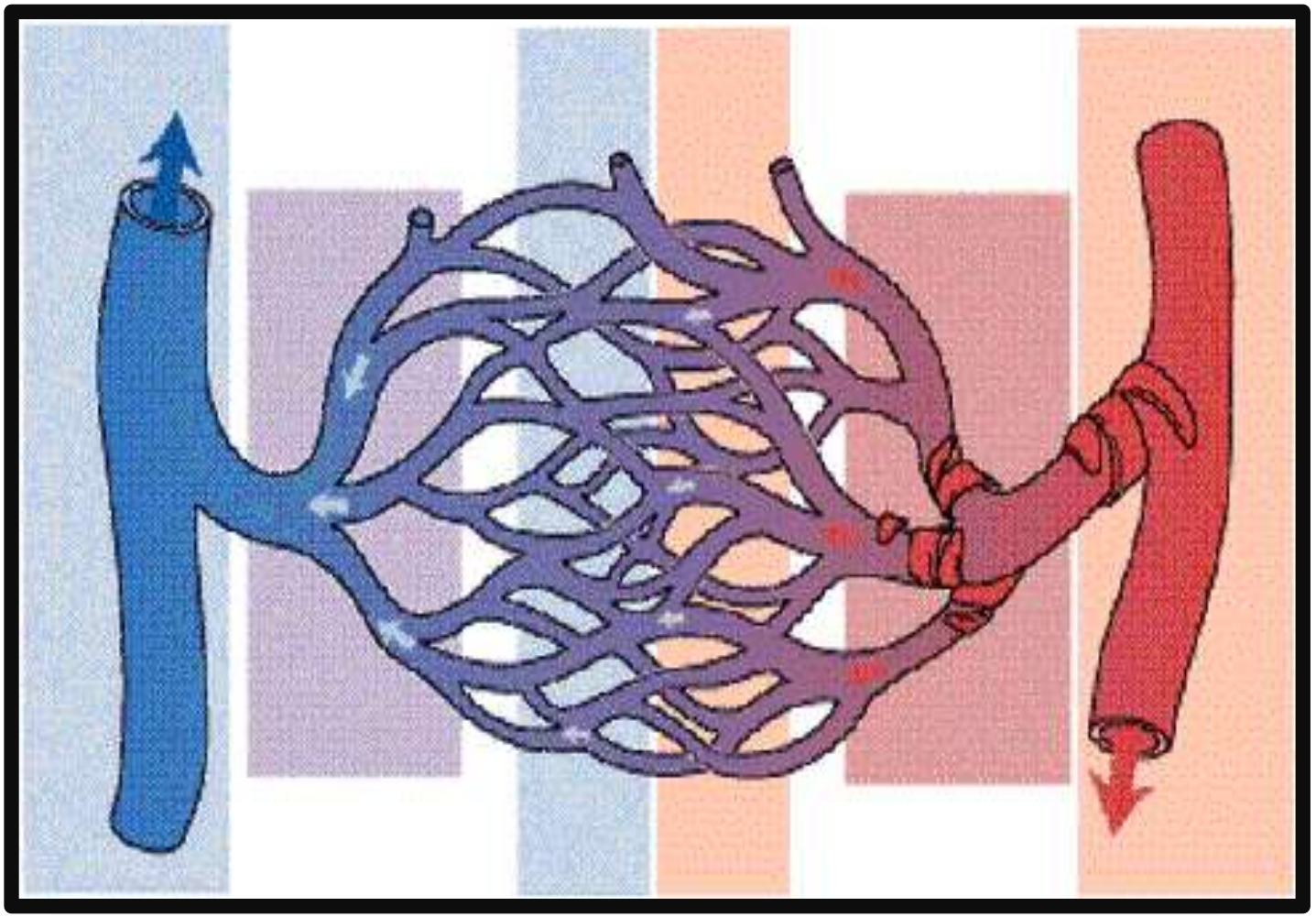
Edema



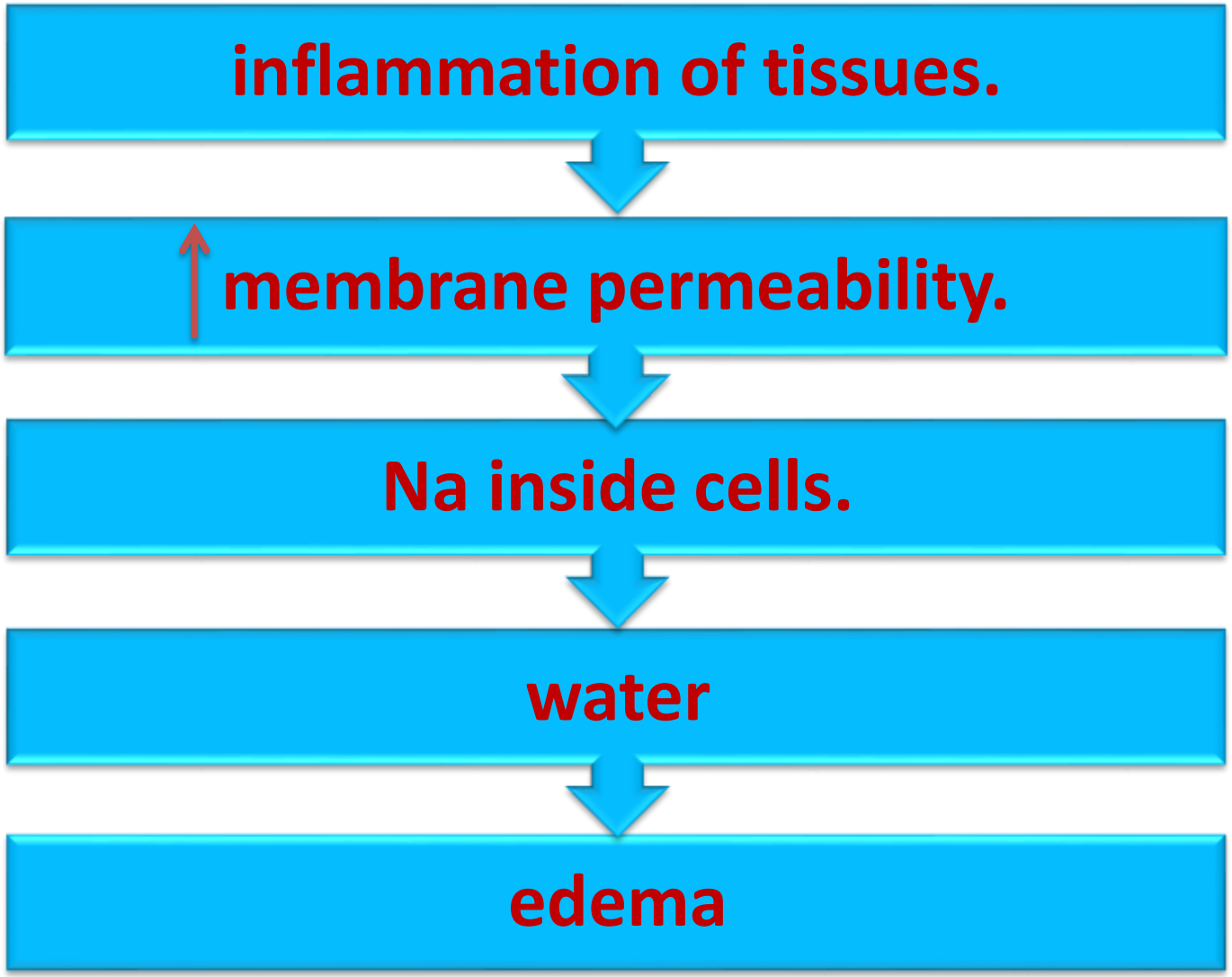
- **Edema occurs mainly in the ECF compartment**

Extracellular Edema

common clinical cause is excessive capillary fluid filtration.



Intracellular Edema:



الله يوفقن